

## Chemistry Student Edition - Basic Answer Key

### Chapter 19: Chemical Equilibrium

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#### The Nature of Chemical Equilibrium

##### Questions

1. At chemical equilibrium, what is the relationship between the rate of the forward reaction and the rate of the reverse reaction?
2. What does the law of mass action say about any reaction at equilibrium?
3. Why are solids and liquids left out of the equilibrium constant expression?
4. Explain why the equilibrium constant expression as derived from concentrations can be converted to the equilibrium constant expression as derived from partial pressures.
5. Construct balanced chemical equations and write the equilibrium constant expression (either  $K_c$  or  $K_p$ ) for each of the following processes.
  - a) Sodium metal reacts vigorously with oxygen gas to produce sodium oxide.
  - b) Aluminum oxide decomposes to form aluminum metal and oxygen gas.
  - c) Nitrogen gas and oxygen gas react to form nitrogen monoxide gas.
6. Give an example of a chemical equilibrium that exists in our daily lives.
7. Write a  $K_{eq}$  expression for each of the following processes.
  - a) Solid calcium carbonate is heated and decomposes into solid calcium oxide and carbon dioxide gas.
  - b) Hydrogen gas ( $H_2$ ) and solid iodine ( $I_2$ ) are combined to form hydrogen iodide gas ( $HI$ ).

##### Answers

1. At equilibrium, the rate of the forward reaction equals the rate of the reverse reaction.
2. The ratio of product concentrations to reactant concentrations (along with the appropriate exponents) is equal to a constant.
3. Solids and liquids have essentially constant densities, which means that writing them is not very useful.
4. Partial pressures are related to concentrations by a factor of  $RT$ .
- 5.



$$K_c = \frac{[Na_2O_{(s)}]^2}{[Na_{(s)}]^2 [O_{2(g)}]}$$

- a) heterogeneous equilibria



$$K_c = \frac{[Al_{(s)}]^4 [O_{2(g)}]^3}{[Al_2O_{3(s)}]^2}$$

- b) heterogeneous equilibria

# Answer Key Ch 19 Chemistry

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*Handbook of sol-gel science and technology. 1. Sol-gel processing* Sumio Sakka, 2005 Since Dr Disiich of Germany prepared a glass lens by the sol gel method around 1970 sol gel science and technology has continued to develop Since then this field has seen remarkable technical developments as well as a broadening of the applications of sol gel science and technology There is a growing need for a comprehensive reference that treats both the fundamentals and the applications and this is the aim of Handbook of Sol Gel Science and Technology The primary purpose of sol gel science and technology is to produce materials active and non active including optical electronic chemical sensor bio and structural materials This means that sol gel science and technology is related to all kinds of manufacturing industries Thus Volume 1 Sol Gel Processing is devoted to general aspects of processing Newly developed materials such as organic inorganic hybrids photonic crystals ferroelectric coatings photocatalysts will be covered Topics in this volume include Volume 2 Characterization of Sol Gel Materials and Products highlights the important fact that useful materials are only produced when characterization is tied to processing Furthermore characterization is essential to the understanding of nanostructured

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