

CHARLES LAW WORKSHEET #2

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

NAME KRY

1. A gas measures 150 mL at 1.0 atmospheres and 20°C. Calculate its volume in mL at 0°C and 1.0 atmospheres.

$$V_1 = 150 \text{ mL}$$

$$T_1 = 20^\circ\text{C} + 273 = 293 \text{ K}$$

$$T_2 = 0^\circ\text{C} + 273 = 273 \text{ K}$$

$$\frac{150 \text{ mL}}{293 \text{ K}} = \frac{V_2}{273 \text{ K}}$$

$$\frac{(150 \text{ mL})(273 \text{ K})}{(293 \text{ K})} = \frac{(293 \text{ K})V_2}{(293 \text{ K})}$$

$$V_2 = 140 \text{ mL}$$

2. A gas occupies a volume of 4.50 liters at 27°C. At what temperature in °C would the volume be 6.0 liters, the pressure remaining constant?

$$V_1 = 4.50 \text{ L} \quad T_1 = 27^\circ\text{C} + 273 = 300 \text{ K}$$

$$V_2 = 6.0 \text{ L} \quad T_2 = ?$$

$$\frac{4.50 \text{ L}}{300 \text{ K}} = \frac{6.0 \text{ L}}{T_2}$$

$$\frac{(4.50 \text{ L})T_2}{4.50 \text{ L}} = \frac{(300 \text{ K})(6.0 \text{ L})}{(4.50 \text{ L})}$$

$$T_2 = 400 \text{ K} - 273 \text{ K}$$

$$= 127^\circ\text{C}$$

3. A sample of gas occupies 220 mL at 10°C and 750 torr. What volume in mL will the gas have at 20°C and 750 torr?

$$V_1 = 220 \text{ mL} \quad T_1 = 10^\circ\text{C} + 273 = 283 \text{ K}$$

$$V_2 = ? \quad T_2 = 20^\circ\text{C} + 273 = 293 \text{ K}$$

$$\frac{220 \text{ mL}}{283 \text{ K}} = \frac{V_2}{293 \text{ K}}$$

$$\frac{(220 \text{ mL})(293 \text{ K})}{283 \text{ K}} = \frac{(283 \text{ K})V_2}{283 \text{ K}}$$

$$V_2 = 228 \text{ mL}$$

4. A gas occupies a volume of 90.0 mL at 27°C and 740 torr. What volume in mL will the gas have at 5°C and 740 torr?

$$V_1 = 90.0 \text{ mL} \quad T_1 = 27^\circ\text{C} + 273 = 300 \text{ K}$$

$$T_2 = 5^\circ\text{C} + 273 = 278 \text{ K}$$

$$V_2 = 83.4 \text{ mL}$$

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \Rightarrow \frac{90.0 \text{ mL}}{300 \text{ K}} = \frac{V_2}{278 \text{ K}}$$

$$\frac{(278 \text{ K})(90.0 \text{ mL})}{(300 \text{ K})} = \frac{V_2(300 \text{ K})}{300 \text{ K}}$$

5. A gas occupies a volume of 120 mL at 27°C and 630 torr. At what temperature in °C would the volume be 80.0 mL at 630 torr?

$$V_1 = 120 \text{ mL} \quad T_1 = 27^\circ\text{C} + 273 = 300 \text{ K}$$

$$V_2 = 80.0 \text{ mL}$$

$$\frac{(120 \text{ mL})T_2}{120 \text{ mL}} = \frac{(80.0 \text{ mL})(300 \text{ K})}{(120 \text{ mL})}$$

$$\frac{120 \text{ mL}}{300 \text{ K}} = \frac{80.0 \text{ mL}}{T_2}$$

$$T_2 = 200 \text{ K} - 273 \text{ K} = -73^\circ\text{C}$$

Answer Key For Boyles Law

ML Morrison

A red circular graphic with a gradient, appearing as a stylized arrow or a decorative element, located to the right of the author's name.

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