

Date _____ Name _____
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Pre-Lab Study Questions | 14

1. Why does an oil-and-vinegar salad dressing have two separate layers?

Salad dressing forms layers because oil is not soluble in water. This causes a heterogeneous mixture, which is seen as two separate layers.

2. What is meant by the mass percent (m/m) concentration of a solution?

It describes the grams of solute in the grams of solution multiplied by 100.

3. What is molarity?

It is a concentration unit. It is the moles of solute in the liters of solution.

4. Why are some electrolytes considered strong, whereas others are considered weak?

**Strong electrolytes, such as strong acids, form many ions in solution.
Weak electrolytes, such as weak acids, form few ions in solution.**

5. A solution is prepared with 3.26 g KCl and water to make 25.0 mL of KCl.
a. What is the % (m/v) of the KCl solution?

$$\%(\text{m}/\text{v})\text{KCl} = \frac{3.26\text{g KCl}}{25.0\text{mL sol'n}} (100) = 13.04 \approx 13.0\%$$

- b. What is the molarity (M) of the KCl solution?

$$M_{\text{KCl}} = \frac{3.26\text{g KCl} \left(\frac{1\text{ mole KCl}}{74.6\text{g KCl}} \right)}{25.0\text{ mL sol'n} \left(\frac{1\text{ L}}{1000\text{ mL}} \right)} = 1.747989 \approx 1.75\text{ M KCl}$$

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