

Acids and bases review sheet Answer key

1. List the differences between an acid and a base. Your description should include such things as whether both are electrolytes, which contains a hydrogen ion and which includes a hydroxide ion, and you should also include two examples of an acid and two examples of a base that are present in your household.

Acid	base
Conducts electricity	conducts electricity
Contains hydrogen ion	contains hydroxide ion
Proton donor	proton acceptor
Turns litmus paper red	turns litmus paper blue
React with most metals	releases hydroxide ion
Releases hydronium	Ex: Windex,
Ex: lemon	bleach
Vinegar	soap
Battery acid (sulfuric acid)	

2. Listed below are some of the properties of acids and bases. Fill in the blanks with the appropriate word, **acids** or **bases**:

- A** produce hydrogen ions (H^+) in solution
- B** have a bitter taste
- B** produce hydroxide ions (OH^-) ions
- A** react with bases to form salts
- A** have a sour taste
- B** have a slippery, soapy feel
- B** react with acids to form salts
- A** react with many metals

3. Write an equation for the dissociation of the following **acids** in water; follow example

- $HClO_4 \rightarrow H^+ + ClO_4^-$
- $H_2SO_4 \rightarrow 2H^+ + SO_4^{2-}$
- $HC_2H_3O_2 \rightarrow H^+ + C_2H_3O_2^-$
- $H_2S \rightarrow 2H^+ + S^{2-}$
- $HCl \rightarrow H^+ + Cl^-$
- $HNO_3 \rightarrow H^+ + NO_3^-$

4. Write an equation for the dissociation (ionization) of the following **bases** in water: (follow example)

- $Mg(OH)_2 \rightarrow Mg^{2+} + 2OH^-$
- $NaOH \rightarrow Na^+ + OH^-$
- $Ba(OH)_2 \rightarrow Ba^{2+} + 2(OH^-)$
- $KOH \rightarrow K^+ + OH^-$
- $Mg(OH)_2 \rightarrow Mg^{2+} + 2(OH^-)$
- $LiOH \rightarrow Li^+ + OH^-$

5) A. What type of reaction is a **neutralization** reaction? **Double replacement**

B. Write the net equation for neutralization:



Answer Key For Review Acids And Bases

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Answer Key For Review Acids And Bases:

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amines and derivatives of guanidine or pyridine The solvents emphasized are hydrocarbons and halohydrocarbons but data for dipolar aprotic solvents for example acetone acetonitrile and nitrobenzene are included Contrasts in acid base behavior and in acidity and basicity scales in aprotic and water like solvents are discussed The role of hydrogen bonding in aprotic solvents is discussed at length Important types of hydrogen bonded structures include chelate rings self associated acids bases and salts hydrogen bonded ion pairs and homo and heteroconjugate cations and anions Examples are given in which hydrogen bonding of these types affects such properties as the absorption spectrum of a salt the catalytic effect of an acid and the accurate location of a titration endpoint Author **Saunders Comprehensive Review for the NCLEX-RN**

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