

Unit 5 Review: Bonding and Metals

Determine which type of bond would occur between the atoms: ionic or covalent.

Atom 1	Atom 2	Type of Bond
Aluminum	Phosphorus	ionic
Sulfur	Oxygen	covalent
Lithium	Nitrogen	ionic
Carbon	Sulfur	covalent

Write the formulas for the following compounds and then determine if the bonds are ionic or covalent. Draw the dot diagram for the compound. If it is covalent, indicate its shape.

Magnesium Oxide
 Type: ionic
 Chemical Formula: MgO
 Dot Diagram:
 $Mg \cdot \cdot \cdot O \cdot \cdot \cdot \rightarrow [Mg]^{2+} [O]^{2-}$
 Shape (if covalent):

Carbon Tetrafluoride
 Type: covalent
 Chemical Formula: CF_4
 Dot Diagram:
 $C \cdot \cdot \cdot F \cdot \cdot \cdot F \cdot \cdot \cdot F \cdot \cdot \cdot F \cdot \cdot \cdot$
 Shape (if covalent): tetrahedral

Methane
 Type: covalent
 Chemical Formula: CH_4
 Dot Diagram:
 $C \cdot \cdot \cdot H \cdot \cdot \cdot H \cdot \cdot \cdot H \cdot \cdot \cdot H \cdot \cdot \cdot$
 Shape (if covalent): tetrahedral

Carbon Monoxide
 Type: covalent
 Chemical Formula: CO
 Dot Diagram:
 $C \cdot \cdot \cdot O \cdot \cdot \cdot$
 Shape (if covalent): linear

Barium Bromide
 Type: ionic
 Chemical Formula: $BaBr_2$
 Dot Diagram:
 $Ba \cdot \cdot \cdot Br \cdot \cdot \cdot Br \cdot \cdot \cdot \rightarrow [Ba]^{2+} [Br]^{1-} [Br]^{1-}$
 Shape (if covalent):

Sodium Oxide
 Type: ionic
 Chemical Formula: Na_2O
 Dot Diagram:
 $Na \cdot \cdot \cdot O \cdot \cdot \cdot Na \cdot \cdot \cdot \rightarrow [Na]^{+} [O]^{2-} [Na]^{+}$
 Shape (if covalent):

Bromic Acid
 Type: covalent
 Chemical Formula: $HBrO_3$
 Dot Diagram:
 $H \cdot \cdot \cdot Br \cdot \cdot \cdot O \cdot \cdot \cdot O \cdot \cdot \cdot O \cdot \cdot \cdot$
 Shape (if covalent): pyramidal

Nitrate Ion
 Type: covalent
 Chemical Formula: NO_3^{-}
 Dot Diagram:
 $N \cdot \cdot \cdot O \cdot \cdot \cdot O \cdot \cdot \cdot$
 Shape (if covalent): trigonal planar

Aluminum Sulfide
 Type: ionic
 Chemical Formula: Al_2S_3
 Dot Diagram:
 $Al \cdot \cdot \cdot S \cdot \cdot \cdot S \cdot \cdot \cdot S \cdot \cdot \cdot S \cdot \cdot \cdot S \cdot \cdot \cdot$
 Shape (if covalent):

Hydrogen Sulfide
 Type: covalent
 Chemical Formula: H_2S
 Dot Diagram:
 $H \cdot \cdot \cdot S \cdot \cdot \cdot H \cdot \cdot \cdot$
 Shape (if covalent): bent

Use the elements in the following list to answer the questions below: Aluminum, Calcium, Copper, Iron, Lead, Magnesium, Mercury, Nickel, Potassium, Silver, Sodium, Tin, Titanium, Zinc

- Give one of the metals from the above list that is in its toxic form is not good for the body.
Mercury
- Give one of the metals from the list above that can be found free in nature.
Silver, Mercury, Copper, Iron
- Give one of the metals from the list above that cannot be found free in nature.
Sodium, Potassium, Magnesium, Calcium, Aluminum, Lead, Nickel, Tin, Titanium
- Name one of the strongest transition metals from the list above.
Titanium, Nickel, Aluminum
- Give one metal from the list above which is found in nature as a chloride.
Sodium, Potassium, Magnesium, Calcium
- Which metal from the list above gets its name from the Latin word maximum?
Tin
- Which metal from the list above gets its name from the Latin word hydrargyrum?
Mercury

Answer the following questions using one of the transition metals that we discussed: Aluminum, Copper, Iron, Lead, Mercury, Nickel, Silver, Tin, Titanium, Zinc

Property	Transition Metal or Metals
Good Conductors (2)	$Ag, Cu, Al, Fe, or Hg$
High Density (1)	$Hg, Pb, or Ag$
Low Density (1)	$Al or Ti$
Soft (2)	$Sn, Pb, or Hg$
Magnetic (1)	$Fe or Ni$
Good Luster (2)	$Ag, Al, Ni, or Hg$

Uses in Pure Form	Transition Metal
Fuel for Space Shuttle	Al
Car Batteries	Pb
Extreme Cold Iron Masks	Hg
Bike Frames	Ti
Wiring	Cu
Inside of Printer	Zn

Group 1 and Group 2 elements have many similarities and differences. In the table below, give a few words about each of the concepts listed.

Property or Idea	Group 1 Elements	Group 2 Elements
Reactivity with water	vigorous	mild
Density	low density	more dense than Grp 1
Occurrence in Nature	never	never
Hardness	soft	harder than Grp 1

Match the location with the metal that we need to get from this area.

Location	Metal or Metals
Bismuthville Plant, Utah	Ni
Epsom, England	Mg
Dead Sea, Palestine	$Na (K, Mg, Ca)$
Morog, Peru	Na
Salt Lake City, Utah	$Na (K, Mg, Ca)$

Fill in the chart below about the four types of bonds we learned about. Depends on electronegativity

Bond Type	Particles	Picture	Force	Melting Point	Conductivity	Other Notes	Examples
Molecular (Covalent)	nonmetal	$H \cdot \cdot \cdot H \cdot \cdot \cdot$	weak	low	no	usually gases	carbon dioxide
Ionic	metal ion nonmetal ion	$Na^+ \cdot \cdot \cdot Cl^-$	strong	very high	no (if ionic)	usually solids	sodium chloride
Network Covalent	nonmetal pattern	$C \cdot \cdot \cdot C \cdot \cdot \cdot C \cdot \cdot \cdot$	strong	very high	no	highly hardest	diamond
Metallic	metal all	$Na \cdot \cdot \cdot Na \cdot \cdot \cdot Na \cdot \cdot \cdot$	strong	varies	yes	ductile malleable	copper

Answer Key Ionic Bonds Review And Reinforce

Bertram G. Katzung, Anthony J. Trevor

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Prentice Hall Science Explorer Michael J. Padilla, 2002 *Structure & Function of the Body - E-Book* Kevin T. Patton, Frank B. Bell, Terry Thompson, Pegg L. Williamson, 2024-06-25

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Masters, Anthony J. Trevor, 2012-03-22 The most up to date comprehensive and authoritative pharmacology text in health medicine Enhanced by more than three hundred illustrations many in full color Organized to reflect the syllabi in many pharmacology courses and in integrated curricula Basic Clinical Pharmacology 12e covers the important concepts students need to know about the science of pharmacology and its application to clinical practice Selection of the subject matter and order of its presentation are based on the authors many years experience in teaching this material to thousands of medical pharmacy dental podiatry nursing and other health science students To be as clinically relevant as possible the book includes sections that specifically address the clinical choice and use of drugs in patients and the monitoring of their effects and case studies that introduce clinical problems in many chapters Presented in full color and enhanced by more than three hundred illustrations Basic Clinical Pharmacology features numerous summary tables and diagrams that encapsulate important information Coverage that spans every aspect of medical pharmacology Basic principles Autonomic drugs Cardiovascular renal drugs Drugs with important actions on smooth muscles Central nervous system drugs Drugs used to treat anemias clotting disorders hyperlipidemia and inflammation and gout Endocrine drugs Chemotherapeutic and immunologic drugs Toxicology Special subjects perinatal geriatric and dermatologic pharmacology Botanical and food supplements and over the counter medications Prescribing Also in this edition Drug Summary Tables conclude most chapters providing a concise summary of the most important drugs General concepts relating to newly discovered receptors receptor mechanisms and drug transporters Descriptions of important new drugs including monoclonal antibodies Papers Presented at the ... Meeting American Chemical Society. Division of Polymer Chemistry, 1990 **Basic & Clinical Pharmacology** Bertram

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