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*транспорта и спецтехники. Том 1* Александр Соловьев, 2015-02-10      The Trademark Register of the United States  
, 1989      **TV Guide** , 1992      **Mediamark Research Apparel & Accessories Report** , 1998      *Желтые страницы*  
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following equation  $2 \text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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 1  $\text{N}_2 + 3 \text{F}_2 \rightarrow 2 \text{NF}_3$   
 2  $2 \text{C}_6\text{H}_{10} + 17 \text{O}_2 \rightarrow 12 \text{CO}_2 + 10 \text{H}_2\text{O}$   
 3  $\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$   
 4  $2 \text{GaBr}_3 + 3 \text{Na}_2\text{SO}_3 \rightarrow \text{Ga}_2\text{SO}_3 + 3 \text{NaBr}$   
 5  $3 \text{SnO}_2 + \text{NF}_3 \rightarrow 3 \text{SnF}_2 + \text{N}_2\text{O}_3$   
 using the following equation  
 $\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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 1  $2 \text{N}_2 + 3 \text{F}_2 \rightarrow 2 \text{NF}_3$   
 2  $\text{CaH}_4 + 17 \text{O}_2 \rightarrow 12 \text{CO}_2 + 10 \text{H}_2\text{O}$   
 3  $\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$   
 4  $2 \text{GaBr}_3 + 3 \text{Na}_2\text{SO}_3 \rightarrow \text{Ga}_2\text{SO}_3 + 3 \text{NaBr}$   
 5  $3 \text{SnO}_2 + \text{NF}_3 \rightarrow 3 \text{SnF}_2 + \text{N}_2\text{O}_3$

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web may 28 2020 write the balanced equation for the reaction and calculate how much  $\text{SiO}_2$  is required to produce 3.00 kg of  $\text{SiC}$   
 $\text{SiO}_2 + 3 \text{C} \rightarrow \text{SiC} + 2 \text{CO}$

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 $\text{N}_2 + \text{F}_2 + \text{NF}_3 + \text{C}_6\text{H}_{10} + 17 \text{O}_2 + \text{CO} + 3 \text{HBr} + \text{KHCO}_3 + \text{H}_2\text{O} + \text{KBr} + 4 \text{GaBr}_3 + \text{Na}_2\text{SO}_3 + \text{Ga}_2\text{SO}_3 + 3 \text{NaBr} + 5 \text{SnO}_2 + \text{NF}_3 + \text{SnF}_2 + \text{N}_2\text{O}_3 + \text{CO}_2 + \text{NaBr}$   
 solve the following stoichiometry grams problems 6

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