

Stoichiometry Lab: Vinegar and Baking Soda

Do Not Write On This Sheet

Purpose: To predict the amount of Carbon Dioxide gas that should be produced in a chemical reaction; then calculate the % yield.



Materials: Baking Soda (NaHCO_3), Vinegar (CH_3COOH), and 2 plastic cups, scale.

Procedure: Find and record the mass of cup A. With cup A still on the scale, add approximately 10.0 g of baking soda to the cup. **The mass does not have to be exact, as long as you carefully record your results.**

Place cup B on the scale, weigh and record approximately 50.0 g of vinegar. **The mass does not have to be exact, as long as you carefully record your results.**

Slowly add vinegar to cup A until the reaction has stopped. **DO NOT** add all of the vinegar, just enough to complete the reaction. Reweigh and record both cup A and B. Calculate the mass of CO_2 that escaped.

Data:

- a. Mass of Cup A _____ g
- b. Mass of Cup A and baking soda _____ g
- c. Calculate mass of baking soda (b - a) _____ g
- d. Mass of Cup B with vinegar _____ g
- e. Mass of Cup B after reaction _____ g
- f. Calculate mass of vinegar poured into Cup A (d - e) _____ g
- g. Mass of Cup A after reaction _____ g
- h. Calculate mass of product after reaction (g - a) _____ g
- i. Calculate baking soda + vinegar (c + f) _____ g
- j. Calculate mass of CO_2 lost (i - h) _____ g

Analysis Of Baking Soda Lab Answers

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Staley, Michael S. Matta, Edward L.
Waterman, Prentice-Hall Staff**

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Merrill Laboratory Chemistry L. Neal Carmichael, David F. Haines, 1990-10 Laboratory Experiments for General, Organic & Biochemistry Frederick A. Bettelheim, Joseph Marvin Landesberg, 1997

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Chemistry John Kenkel, 2000-06-23 Have you ever had a discussion with an industrial chemist about the job Have you ever shadowed a chemist or chemical technician in an industrial or government laboratory for a day If you have done these things you were likely surprised at how foreign the language seemed or startled at how unfamiliar the surroundings were Was there any talk of t

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