

Calculations in Analytical Chemistry

The Mole

The mole is the SI unit for amount of substance. It corresponds to the amount of a chemical substance that contains the same number of particles present in 12 grams of C-12 which is equivalent to Avogadro's number, that is, $N_A = 6.022 \times 10^{23}$.

The molar mass is given as the mass of a compound in grams per mole of that substance. It is calculated by adding the molar masses of all the component elements in the compound. Take glucose for example:

Element	Molar mass per element	Number of atoms	Mass
C	12 g/mol	6	72 g/mol
H	1 g/mol	12	12 g/mol
O	16 g/mol	6	96 g/mol
Molar Mass			180 g/mol

Mole and Millimole Calculations

The millimole is simply 1/1000 of a mole. It is sometimes used in calculations because of its convenience.

Example (Skoog, et al., 2014):

1. Find the number of moles and millimoles of benzoic acid (122.1 g/mol) that are contained in 2.00 g of pure acid.

If we use HBz to represent benzoic acid:

$$HBz_{mol} = 2.00 \text{ g}_{HBz} \times \frac{1 \text{ mol}_{HBz}}{122.1 \text{ g}_{HBz}} = 0.0164 \text{ mol}_{HBz}$$

To obtain the number of millimoles, the millimolar mass of benzoic acid will be used:

$$HBz_{mol} = 2.00 \text{ g}_{HBz} \times \frac{1 \text{ mmol}_{HBz}}{0.1221 \text{ g}_{HBz}} = 16.4 \text{ mmol}_{HBz}$$

2. What is the mass in grams of Na^+ (22.99 g/mol) in 25.0 g of Na_2SO_4 (142.0 g/mol)?

$$\begin{aligned} \text{mass}_{\text{Na}^+} &= 25.0 \text{ g}_{\text{Na}_2\text{SO}_4} \times \frac{1 \text{ mol}_{\text{Na}_2\text{SO}_4}}{142.0 \text{ g}_{\text{Na}_2\text{SO}_4}} \times \frac{2 \text{ mol}_{\text{Na}^+}}{1 \text{ mol}_{\text{Na}_2\text{SO}_4}} \times \frac{22.99 \text{ g}_{\text{Na}^+}}{1 \text{ mol}_{\text{Na}^+}} \\ \text{mass}_{\text{Na}^+} &= 8.10 \text{ g}_{\text{Na}^+} \end{aligned}$$

Analytical Chemistry Calculation Questions And Answers

Peter C. Meier, Richard E. Zünd



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Table of Contents Analytical Chemistry Calculation Questions And Answers

1. Understanding the eBook Analytical Chemistry Calculation Questions And Answers
 - The Rise of Digital Reading Analytical Chemistry Calculation Questions And Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Analytical Chemistry Calculation Questions And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Analytical Chemistry Calculation Questions And Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Analytical Chemistry Calculation Questions And Answers
 - Personalized Recommendations
 - Analytical Chemistry Calculation Questions And Answers User Reviews and Ratings

- Analytical Chemistry Calculation Questions And Answers and Bestseller Lists
- 5. Accessing Analytical Chemistry Calculation Questions And Answers Free and Paid eBooks
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 - Analytical Chemistry Calculation Questions And Answers Budget-Friendly Options
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 - Analytical Chemistry Calculation Questions And Answers Enhanced eBook Features
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 - Highlighting and Note-Taking Analytical Chemistry Calculation Questions And Answers
 - Interactive Elements Analytical Chemistry Calculation Questions And Answers
- 8. Staying Engaged with Analytical Chemistry Calculation Questions And Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Analytical Chemistry Calculation Questions And Answers
- 9. Balancing eBooks and Physical Books Analytical Chemistry Calculation Questions And Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Analytical Chemistry Calculation Questions And Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Analytical Chemistry Calculation Questions And Answers
 - Setting Reading Goals Analytical Chemistry Calculation Questions And Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Analytical Chemistry Calculation Questions And Answers
 - Fact-Checking eBook Content of Analytical Chemistry Calculation Questions And Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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