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2008 Ford Edge Shop Manual

Phil Edmonston



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end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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 $HCl(aq) + NaOH(aq) \rightarrow NaCl(aq) + H_2O(l)$
 one mole of HCl would be fully neutralized by one mole of NaOH

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solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh reaction with a 2nd substance H_2CO_3 aq 2NaOH aq Na_2CO_3 $4\text{H}_2\text{O}$ to differentiate an acid from a base in a solution a ph scale is used the ph scale ranges from 0 to 14 a ph of 7 is neutral in an acid base reaction

conclusion titration lab - Jan 29 2022

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data calculations and conclusion acid base titration lab - Jun 02 2022

web conclusion in this lab we used titration to explore the concepts of stoichiometry and equivalence points we found the concentration of an unknown substance by mixing 2m hcl with the naoh of unknown concentration in order to experimentally ascertain the concentration of the naoh

this is a chemistry lab report on an acid base titration experiment - Dec 28 2021

web this point of titration is referred to as the equivalence point considering the 1 1 stoichiometry of this acid base reaction NaOH aq $\text{C}_6\text{H}_4\text{COOH}$ aq $\text{C}_6\text{H}_4\text{COONa}$ aq H_2O l the point of equivalence is the point of titration when the number of moles of naoh added is equal to the number of moles of khp in the