

Chemistry

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(Chapter 1)(Solutions) (Intext Questions)

(Class 12)

Question 1.1:

Calculate the mass percentage of benzene (C_6H_6) and carbon tetrachloride (CCl_4) if 22 g of benzene is dissolved in 122 g of carbon tetrachloride.

Answer 1.1:

$$\begin{aligned}\text{Mass percentage of } C_6H_6 &= \frac{\text{Mass of } C_6H_6}{\text{Total mass of the solutions}} \times 100 \% \\ &= \frac{\text{Mass of } C_6H_6}{\text{Mass of } C_6H_6 + \text{Mass of } CCl_4} \times 100 \% = \frac{22}{22 + 122} \times 100 \% = 15.28 \%\end{aligned}$$

$$\begin{aligned}\text{Mass percentage of } CCl_4 &= \frac{\text{Mass of } CCl_4}{\text{Total mass of the solutions}} \times 100 \% \\ &= \frac{\text{Mass of } CCl_4}{\text{Mass of } C_6H_6 + \text{Mass of } CCl_4} \times 100 \% = \frac{122}{22 + 122} \times 100 \% = 84.72 \%\end{aligned}$$

Alternatively, Mass percentage of CCl_4 = (100 - 15.28) % = 84.72%

Question 1.2:

Calculate the mole fraction of benzene in solution containing 30% by mass in carbon tetrachloride.

Answer 1.2:

Let the total mass of the solution be 100 g and the mass of benzene be 30 g.

∴ Mass of carbon tetrachloride = (100 - 30) g = 70 g

Molar mass of benzene (C_6H_6) = (6 × 12 + 6 × 1) g mol⁻¹ = 78 g mol⁻¹

∴ Number of moles of C_6H_6 = $\frac{30}{78} \text{ mol} = 0.3846 \text{ mol}$

Molar mass of carbon tetrachloride (CCl_4) = 1 × 12 + 4 × 35.5
= 154 g mol⁻¹

∴ Number of moles of CCl_4 = $\frac{70}{154} \text{ mol} = 0.4545 \text{ mol}$

Thus, the mole fraction of C_6H_6 is given as:

$$\begin{aligned}&= \frac{\text{Number of moles of } C_6H_6}{\text{Number of moles of } C_6H_6 + \text{Number of moles of } CCl_4} \\ &= \frac{0.3846}{0.3846 + 0.4545} \\ &= 0.458\end{aligned}$$

Question 1.3:

Calculate the molarity of each of the following solutions:

(a) 30 g of $Co(NO_3)_2 \cdot 6H_2O$ in 4.3 L of solution

(b) 30 mL of 0.5 M H_2SO_4 diluted to 500 mL

Answer 1.3:

Molarity is given by: $\text{Molarity} = \frac{\text{Moles of solute}}{\text{Volume of solution in litre}}$

(a) Molar mass of $Co(NO_3)_2 \cdot 6H_2O$ = 59 + 2 (14 + 3 × 16) + 6 × 18 = 291 g mol⁻¹

∴ Moles of $Co(NO_3)_2 \cdot 6H_2O$ = $\frac{30}{291} \text{ mol} = 0.103 \text{ mol}$

Therefore, molarity = $\frac{0.103 \text{ mol}}{4.3 \text{ L}} = 0.023 \text{ M}$

(b) Number of moles present in 1000 mL of 0.5 M H_2SO_4 = 0.5 mol

∴ Number of moles present in 30 mL of 0.5 M H_2SO_4 = $\frac{0.5 \times 30}{1000} \text{ mol} = 0.015 \text{ mol}$

Therefore, molarity = $\frac{0.015 \text{ mol}}{0.5 \text{ L}} = 0.03 \text{ M}$

8 Chemistry Sol Answers

**William Swan Stallybrass (formerly
Sonnenschein.), William Swan
Sonnenschein**



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Solutions Manual to Accompany Elements of Physical Chemistry David Smith, 2013-05-30 The Solutions Manual to accompany Elements of Physical Chemistry 6th edition contains full worked solutions to all end of chapter discussion questions and exercises featured in the book The manual provides helpful comments and friendly advice to aid understanding It is also a valuable resource for any lecturer who wishes to use the extensive selection of exercises featured in the text to support either formative or summative assessment and wants labour saving ready access to the full solutions to these questions *Solutions Manual to Accompany Physical Chemistry for the Life Sciences* C. A. Trapp, M. P. Cady, 2011 This solutions manual contains fully worked solutions to all end of chapter discussion questions and exercises featured in Physical Chemistry for the Life Sciences Longman Science Chemistry 9 Kohli Nitin, 2008-09 **Excel With Objective Chemistry For Iit-Screening** Prof. S. K. Khanna, Dr. N. K. Sharma, Dr. B. Kapila, 2006 **Chapter-wise NCERT + Exemplar + Practice Questions with Solutions for CBSE Chemistry Class 11 - 2nd Edition** Disha Experts, 2017-08-29 The book Chapter wise NCERT Exemplar Practice Questions with Solutions for CBSE Class 11 Chemistry has been divided into 3 parts Part A provides detailed solutions Question by Question of all the questions exercises provided in the NCERT Textbook Part B provides solutions to the questions in the NCERT Exemplar book Part C provides selected Practice Questions useful for the Class 11 examination along with detailed solutions The solutions have been designed in such a manner Step by Step that it would bring 100% Concept Clarity for the student Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction *The Best Books* William Swan Sonnenschein, 1895 **JAM IIT Chemistry Question Bank MSC Entrance** Mocktime Publication, The IIT JAM Chemistry code CY question bank includes topic wise drills Inorganic Organic Physical Chemistry Analytical Environmental Chemistry previous year JAM papers in PDF with solutions books with aligned notes and model test series The exam pattern for JAM Chemistry consists of 60 questions 100 marks across Sections A B C MCQs MSQs NATs consistent with IIT JAM structure The question bank aids entrants by exposing them to high frequency topics helping develop speed and accuracy and offering revision material and full length tests that mirror actual exam conditions Ideal for MSc entrance preparation in top institutes *The Chemical News and Journal of Physical Science*, 1893 *A Reader's Guide to the Choice of the Best Available Books (about 50,000) in Every Department of Science, Art & Literature, with the Dates of the First & Last Editions, & the Price, Size & Publisher's Name of Each Book* William Swan Sonnenschein, 1901 **A Reader's Guide to Contemporary Literature** William Swan Sonnenschein, 1895 **33 Years Chapterwise Solutions NEET**

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