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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

| bond                  | kJ mol <sup>-1</sup> | kJ mol <sup>-1</sup> |
|-----------------------|----------------------|----------------------|
| bond energy in kJ/mol | 440                  | 435                  |

Use the bond energies in Table 3.1 and  $\Delta H$  to calculate the bond energy of one C-H bond in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.
 
$$\begin{aligned}
 & (500) + 3(440) \\
 & 440 + 3(440) \quad \quad \quad 2160 \text{ kJ}
 \end{aligned}$$
- Calculate the energy released when bonds form in the products.
 
$$\begin{aligned}
 & \text{bond broken} - \text{bond formed} = -90 \\
 & 2160 - x = -90 \\
 & x = 2250 \text{ kJ}
 \end{aligned}$$
- Calculate the energy of one C-H bond.
 
$$\begin{aligned}
 & 6 \times \text{C-H} = 2160 \\
 & \text{C-H} = \frac{2160}{6} \quad \quad \quad 360 \text{ kJ/mol}
 \end{aligned}$$

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**Roman Wölfel**



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