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IGCSE A* MODEL ANSWER

0620 CHEMISTRY

May/June 2024

Paper 42



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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

Bond	Energy	Unit
N-H	400	kJ/mol
S-O	500	kJ/mol
O-H	460	kJ/mol
N-O	300	kJ/mol
C-H	410	kJ/mol

Use the bond energies in Table 3.1 and $\Delta H = -204 \text{ kJ/mol}$ to calculate the bond energy of an N-O bond in kJ/mol .

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.

$$\begin{aligned}
 &4(\text{N-H}) + 1(\text{H-H}) \\
 &4(400) + 1(436) \\
 &2160 \text{ kJ}
 \end{aligned}$$
- Calculate the energy released when bonds form in the products.

$$\begin{aligned}
 &\Delta H = \text{Energy needed to break bonds} - \text{Energy released when bonds form} \\
 &-204 = 2160 - x \\
 &x = 2160 + 204 \\
 &x = 2364 \text{ kJ}
 \end{aligned}$$
- Calculate the energy of an N-O bond.

$$\begin{aligned}
 &6 \times (\text{N-O}) = 2364 \\
 &\text{N-O} = \frac{2364}{6} \\
 &394 \text{ kJ/mol}
 \end{aligned}$$

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VM Jensen



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