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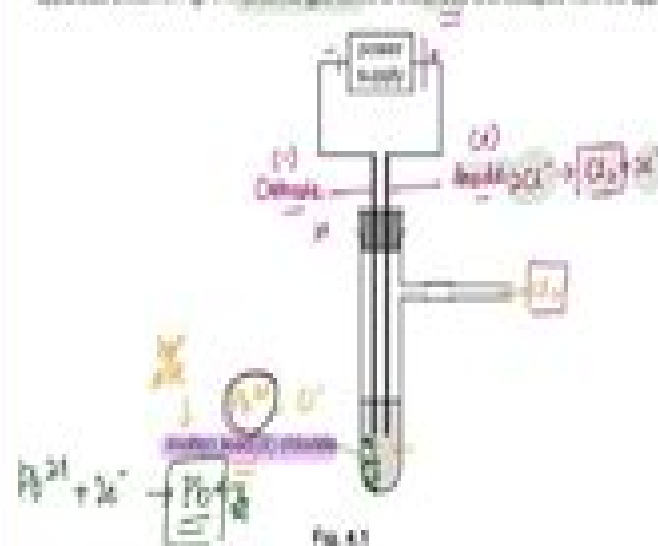
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(a) The student carries out an electrolysis experiment on molten lead(II) chloride using the apparatus shown in Fig. 4.1. (3) (Total gas taken at the anode and escapes from the apparatus.)



(i) Explain why lead(II) chloride needs to be molten before it will conduct electricity / react.
first moving electrons / mobile ions

(ii) Write the ionic half equation for the reaction occurring at the anode.



(iii) State the test for chlorine gas and describe the yellow fumes.

test Damp litmus paper

observations bleached / turns white

(iv) Describe what is observed at the cathode.

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A Loxley



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