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IGCSE A* MODEL ANSWER

0620 CHEMISTRY

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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

bond	H-H	N-N
bond energy in kJ/mol	435	193

Use the bond energies in Table 3.1 and $\Delta H = -91 \text{ kJ/mol}$ to calculate the bond energy of an N-H bond in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.

$$\begin{aligned}
 & (4 \text{ H-H}) + 1 (\text{N-N}) \\
 & 4(435) + 1(193) \quad \quad \quad 2150 \text{ kJ}
 \end{aligned}$$
- Calculate the energy released when bonds form in the products.

$$\begin{aligned}
 & \text{bond broken} - \text{bond formed} = -91 \\
 & 2150 - x = -91 \\
 & x = 2150 + 91 \\
 & x = 2241 \text{ kJ}
 \end{aligned}$$
- Calculate the energy of per mole bond.

$$\begin{aligned}
 & 6 \times (\text{N-H}) = 2241 \\
 & \text{N-H} = \frac{2241}{6} \quad \quad \quad 373.5 \text{ kJ/mol}
 \end{aligned}$$

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R Bogdan



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