

Chapter 10 Chemical Quantities Review

Part 1: Vocabulary

Label the following descriptions with the correct vocabulary word.

- | | |
|---------------------------|--|
| _____ mass | _____ percent composition |
| _____ empirical formula | _____ molar mass |
| _____ number of particles | _____ standard temperature and pressure (STP, 1 atm) |

1. _____ is a description of the relative amounts of each element in a compound.
2. _____ is the lowest whole-number ratio of the atoms of the elements in a compound.
3. _____ is the lowest structural formula and mass of any gas at STP.
4. _____ is the lowest structural formula of a compound.
5. _____ is the lowest temperature of 0°C in 273°K temperature particles of a substance.
6. _____ is 6.02×10^{23} particles.
7. _____ is the lowest mass of a compound at STP. For mass of gas samples is calculated (STP).

Part 2: Section Concept Review

10.1 The Mole: A Measurement

Complete the following statement to demonstrate knowledge of the terms and your understanding of the concepts involved with the terms of mole measurement concepts.

Knowing what mole counting, atoms, and volume for a single quantity called the _____ number. _____ represents the number of particles in a sample of a substance is _____ Avogadro's number.

By knowing the mass of a sample of a compound, scientists use together the _____ number, _____ of the substance and the percentage, other use techniques the sample the mass (mole) the _____ number _____ of the compound. These are _____ 6.02×10^{23} _____ represents the number of particles.

True or False

_____ 1. A mole of any substance has a mass of 1 gram.

_____ 2. A mole of substance contains 6.02×10^{23} atoms.

_____ 3. The temperature particles of a substance is the temperature.

_____ 4. A mole of H_2O is composed of one atom of hydrogen and two atoms of oxygen.

_____ 5. The mass mass of a sample gas is the _____.

Problems

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