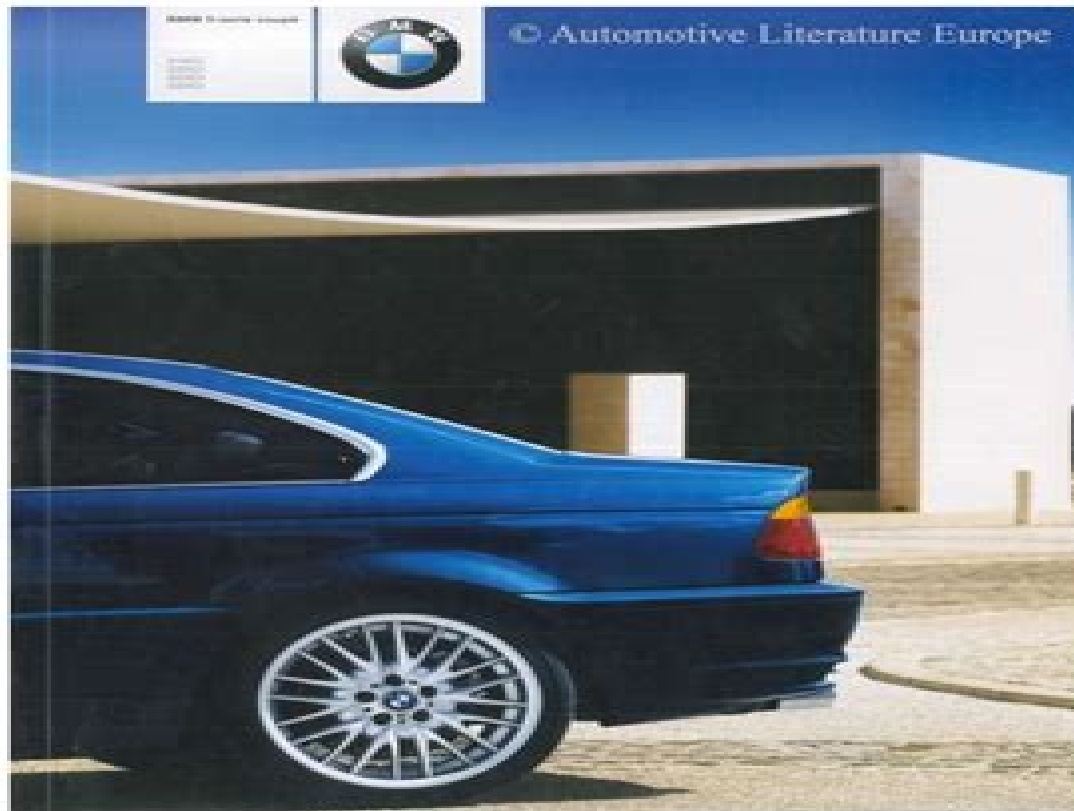


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2001 Bmw 3 Series Coupe Owners Manual

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web a reaction and its experimental rate law are represented below $2 \text{ICl} + \text{H}_2 \text{g} \rightarrow 2 \text{HCl} + \text{I}_2 \text{g}$ rate $k [\text{ICl}]^2 [\text{H}_2]$ a chemist proposes a mechanism for the reaction that is consistent with the rate law the mechanism has two elementary steps and the first step is slow compared to the second

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 $\text{O}_3 \rightarrow \text{O}_2 + \text{O}$ fast
 $\text{O} + \text{O}_3 \rightarrow \text{O}_2 + \text{O}_2$ slow
 a write the overall balanced chemical equation
 $2 \text{O}_3 \rightarrow 3 \text{O}_2$
 b identify any intermediates within the mechanism
 c what is the order with respect

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consider the following mechanism
 $\text{O}_3 \xrightarrow{\text{fast}} \text{O}_2 + \text{O} \xrightarrow{\text{slow}} \text{O}_2 + \text{O}_2$
 a write the overall balanced chemical equation
 $2 \text{O}_3 \rightarrow 3 \text{O}_2$
 b identify any intermediates within the mechanism

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 $\text{Mg} + 2 \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$
 the reaction between Mg and HCl is represented by the equation above in a kinetics experiment
 a 0.080 g sample of Mg ribbon is added to an erlenmeyer flask containing 50 ml of 3.0 M HCl at 25 °C

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