

Empirical and Molecular Formula Practice

3 pages + 3-page Answer KEY + Periodic Table

EMPIRICAL AND MOLECULAR PROBLEMS

Chemistry

1 The percent composition of a certain gas is nitrogen (80.0%) and hydrogen (20.0%). Calculate the empirical formula.

Step 1. Convert % to g.

N = 80

H = 20

Step 2. Convert g to mol.

Step 3. Divide by the smallest mole value.

Step 4. Multiply by smallest number that will make a whole number.

2 The compound in problem 1 has a molar mass of 16 g/mol. Calculate the molecular formula.

Step 1. Calculate the molar mass of the empirical formula.

Step 2. Divide molecular molar mass by empirical molar mass.

Step 3. Multiply empirical formula by the answer in Step 2.

3 A compound is composed of 82.7% nitrogen and 17.3% hydrogen. What is the empirical formula of this compound?

4 The compound in problem 3 has a molar mass of 16 g/mol. Calculate the molecular formula.

5 Hydrazine, a high energy fuel, is 87.5% nitrogen and 12.5% hydrogen. Calculate the empirical formula.

6 Hydrazine's molar mass is 32 g/mol. Calculate the molecular formula.

7 A compound, composed of hydrogen (5.9%) and oxygen (94.1%), has a molar mass of 34 g/mol. What is the empirical formula of this compound?

8 What is the molecular formula of the compound in problem 7?

9 A compound with a known molar mass of 28 g/mol contains carbon (85.7%) and hydrogen (14.3%). What is the empirical formula?

10 What is the molecular formula of the compound in problem 9?

ANSWER KEY INCLUDED

Answer Key Empirical And Molecular Practice

D Kirk



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