

2. Chapter 1: Fundamental Structures & Kinetic Significance of Enzymes

Question 1.1: Structure and Thermodynamics

- a) Which groups lead to the fact that the thermodynamic stability of H_2O , CO_2 , and nucleic acids, CH_2O (sugar's precursor), NH_3 and H_2 is 0?
- b) We speak of enthalpy
- c) Enthalpy is not a state function, but a path-dependent quantity for any system.
- d) At standard conditions, enthalpies of formation and Gibbs free energy of formation values suggest spontaneous reactions. Thermodynamic states suggest reactions from left to right.
- e) Free Enthalpy (Gibbs free energy) of a free thermodynamic system is the maximum work that can be obtained from it.
- f) If the system is not in equilibrium with its surroundings, the thermodynamic state is called non-equilibrium.
- g) At work of thermodynamic processes, ΔG is negative.
- h) The Gibbs free energy is the sum of enthalpy and entropy, and the entropy is the sum of enthalpy and entropy.
- i) ΔG is the free energy of a system. It is the sum of enthalpy and entropy.

Question 1.2: Kinetic Significance of Enzymes (Kinetic Significance)

- a) Enthalpy is not a state function, but a path-dependent quantity for any system.
- b) Enthalpy is not a state function, but a path-dependent quantity for any system.
- c) Enthalpy is not a state function, but a path-dependent quantity for any system.
- d) If the system is not in equilibrium with its surroundings, the thermodynamic state is called non-equilibrium.
- e) At equilibrium, the Gibbs free energy is the sum of enthalpy and entropy.
- f) If the system is not in equilibrium with its surroundings, the thermodynamic state is called non-equilibrium.
- g) If the system is not in equilibrium with its surroundings, the thermodynamic state is called non-equilibrium.
- h) If the system is not in equilibrium with its surroundings, the thermodynamic state is called non-equilibrium.
- i) If the system is not in equilibrium with its surroundings, the thermodynamic state is called non-equilibrium.

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