

Qualitative Analysis and Chemical Bonding

Concepts:

- Chemical bonding
- Covalent bonding
- Ionic bonding
- Metallic bonding
- Intermolecular bonding
- Physical & chemical properties
- Electronegativity

Introduction

- Types of atoms are not equal in chemical bonds: **electronegativity**
- In order to understand how all substances behave we need to know how they are bonded, strong
- In order to know a substance behaves how strongly, strong, weak, we need to know the nature of the chemical bond

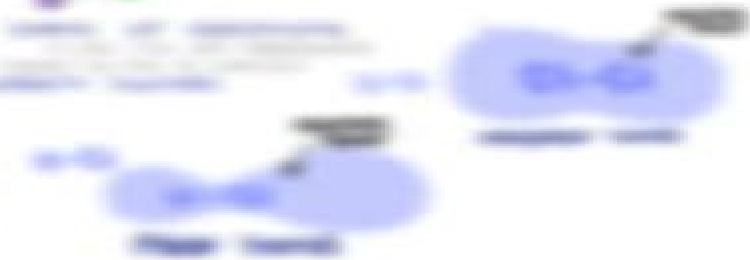
Electronegativity

- Electronegativity: tendency of an atom to attract electrons towards itself in a covalent compound
- More electronegative atom is more likely to attract electrons in a covalent bond, more electronegative, stronger bond
- Electronegativity increases with the strength of a bond
- More electronegative atoms attract electrons more strongly, more electronegative, stronger bond



Covalent bonding

Two atoms share electrons to form a covalent bond. The shared electrons are represented by a pair of dots between the two atoms.

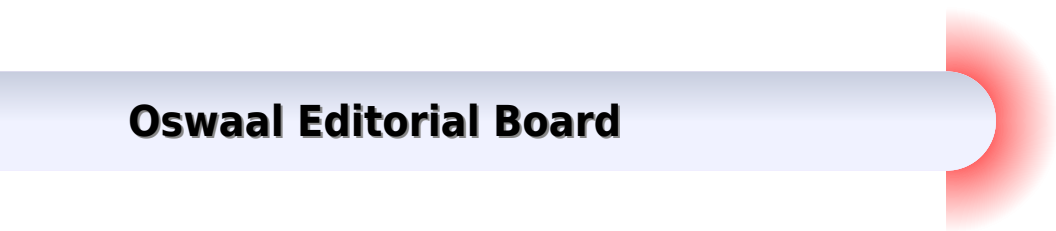


Electronegativity is a measure of an atom's ability to attract electrons towards itself in a covalent bond. The more electronegative an atom is, the more it will attract electrons.

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