

Bachelor of Engineering Examination, 2023

(1st Year, 1st Semester)

CHEMISTRY

Time: 3 hours

Full Marks: 100

Use separate answer sheet for Chemistry-I, Chemistry-II and Chemistry-III

CO-1

CHEMISTRY-I

Answer any three questions from CO-1

Q1. (a) Derive the expression for the calculation of the radius of Bohr's orbit and the velocity of the electron in hydrogen atom. 2+2

(b) What would be the wavelength of the 2nd line in the Lyman series for the H atom [Given $R_H = 109678 \text{ cm}^{-1}$]. 2

(c) Calculate the de Broglie wavelength of a bullet having mass = $2 \times 10^{-3} \text{ kg}$, moving with the speed of 300 ms^{-1} . 2

(d) State and explain "Heisenberg uncertainty principle". 2

Q2. (a) What are the merits and demerits of Bohr's atomic model? 3

(b) Calculate the ionization energy of Li^{2+} ion. 1.5

(c) Define lattice energy. Construct Born-Haber cycle of NaCl. 1 + 2.5

(d) Applying Fajan's rule explain the trend on thermal stabilities of the following compounds

. BeCO_3 , MgCO_3 , CaCO_3 , SrCO_3 , BaCO_3 . 2

Q3. (a) Define buffer solutions with examples. 2

(b) Discuss the mechanism of buffer action. 3

(c) Determine the pH of a solution obtained by mixing equal volumes of 0.1 M NH_4Cl and 0.02 M NH_4OH . [Given that K_b for NH_4OH is 1.8×10^{-5}] 5

Q3. (a) Explain the trend of H-X-H angles (X = central atom) in CH_4 , NH_3 and H_2O .

3

[Turn over

Answer To 1st Semester Chemistry Exam Review

Neil D. Jespersen, Pamela Kerrigan



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Phlebotomy Exam Review, Enhanced Edition Ruth E. McCall, 2020-06-19 Phlebotomy Exam Review Enhanced Seventh Edition thoroughly prepares students for any of the national certification exams in phlebotomy

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