

Name _____ Date _____ Class _____

CHAPTER 15 **STUDY GUIDE FOR CONTENT MASTERY**

Solutions

Section 15.1 What are solutions?

In your textbook, read about the characteristics of solutions.

Use each of the terms below just once to complete the passage.

immiscible	liquid	soluble	solution
insoluble	miscible	solvent	solvent

Air is air! (1) solution of oxygen gas dissolved in nitrogen gas. The oxygen is air is the (2) solvent and nitrogen is the (3) solvent. Because oxygen gas dissolves in a solvent, oxygen gas is air! (4) soluble substances. A substance that does not dissolve is (5) insoluble. (6) liquid solutions are the most common type of solutions. If one liquid is soluble in another liquid, such as acetic acid in water, the two liquids are (7) miscible. However, if one liquid is insoluble in another, the liquids are (8) immiscible.

Read about solvation in aqueous solutions in your textbook.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution. Use the diagram to answer the following questions.



8. Hydration is solvation in which the solvent is water. What is solvation?

Solvation is the process of surrounding solute particles with solvent particles to form a solution.

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Section 15.1 continued

10. As sodium chloride dissolves in water, what happens to the sodium and chloride ions?

The sodium and chloride ions are separated and surrounded by the water molecules.

11. Explain the attraction of the water molecules around the sodium ions and chloride ions.

Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atoms). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

12. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solvation process to occur.

13. List three ways that the rate of solvation may be increased.

stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent

In your textbook, read about heat of solution, solubility, and factors that affect solubility.

For each statement below, write true or false.

True _____ 14. The overall energy change that occurs when a solution forms is called the heat of solution.

False _____ 15. Solubility is a measure of the maximum amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.

False _____ 16. Solvation continues as long as the solvation rate is less than the crystallization rate.

True _____ 17. In a saturated solution, solvation and crystallization are in equilibrium.

True _____ 18. Additional solute can be dissolved in an unsaturated solution.

True _____ 19. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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**Donald Voet, Judith G. Voet, Charlotte
W. Pratt**



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Fundamentals of Biochemistry Donald Voet, Judith G. Voet, Charlotte W. Pratt, 2016-02-29 Voet Voet and Pratt's Fundamentals of Biochemistry 5th Edition addresses the enormous advances in biochemistry particularly in the areas of structural biology and Bioinformatics by providing a solid biochemical foundation that is rooted in chemistry to prepare students for the scientific challenges of the future While continuing in its tradition of presenting complete and balanced coverage that is clearly written and relevant to human health and disease Fundamentals of Biochemistry 5e includes new pedagogy and enhanced visuals that provide a pathway for student learning

Resources in Education, 1995 The Impact of the 4th Industrial Revolution on Engineering Education Michael E. Auer, Hanno Hortsch, Panarit Sethakul, 2020-03-17 This book gathers papers presented at the 22nd International Conference on Interactive Collaborative Learning ICL2019 which was held in Bangkok Thailand from 25 to 27 September 2019 Covering various fields of e learning and distance learning course and curriculum development knowledge management and learning real world learning experiences evaluation and outcomes assessment computer aided language learning vocational education development and technical teacher training the contributions focus on innovative ways in which higher education can respond to the real world challenges related to the current transformation in the development of education Since it was established in 1998 the ICL conference has been devoted to new approaches in learning with a focus on collaborative learning Today it is a forum for sharing trends and research findings as well as presenting practical experiences in learning and engineering pedagogy The book appeals to policymakers academics educators researchers in pedagogy and learning theory school teachers and other professionals in the learning industry and further and continuing education

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