

## Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- Write the balanced chemical equation
- Identify the given (with units) and what you want to find (with units)
- Show set up with units. Check sig figs, give final answer with units and label.

1. Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?



Before  $\underline{2 \text{ mol}}$   $\underline{0 \text{ mol}}$   $\underline{0 \text{ mol}}$

Change  $\underline{-2 \text{ mol}}$   $\underline{+2 \text{ mol}}$  \_\_\_\_\_

After  $\underline{0 \text{ mol}}$   $\underline{2 \text{ mol}}$

$$36 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.0 \text{ g H}_2\text{O}} = 2.0 \text{ mol H}_2\text{O} \times \frac{2 \text{ mol H}_2}{2 \text{ mol H}_2\text{O}} = 2 \text{ mol H}_2 \times \frac{2.0 \text{ g H}_2}{1 \text{ mol H}_2} = 4.0 \text{ g H}_2$$

2. Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?



Before  $\underline{9.94 \text{ mol}}$   $\underline{\text{xs mol}}$   $\underline{0 \text{ mol}}$

Change  $\underline{-9.94 \text{ mol}}$  \_\_\_\_\_  $\underline{+9.94 \text{ mol}}$

After  $\underline{0 \text{ mol}}$   $\underline{9.94 \text{ mol}}$

$$581 \text{ g NaCl} \times \frac{1 \text{ mol NaCl}}{58.5 \text{ g NaCl}} = 9.93 \text{ mol NaCl} \times \frac{2 \text{ mol Na}}{2 \text{ mol NaCl}} = 9.93 \text{ mol Na} \times \frac{23.0 \text{ g Na}}{1 \text{ mol Na}} = 228 \text{ g Na}$$

3. You eat 180.0 g of glucose (90 M&Ms). If glucose,  $\text{C}_6\text{H}_{12}\text{O}_6$ , reacts with oxygen gas to produce carbon dioxide and water, how many grams of oxygen will you have to breathe in to burn the glucose?



Before  $\underline{1.0 \text{ mol}}$   $\underline{\text{xs mol}}$   $\underline{0 \text{ mol}}$   $\underline{0 \text{ mol}}$

Change  $\underline{-1.0 \text{ mol}}$  \_\_\_\_\_  $\underline{-1.0 \text{ mol}}$

After  $\underline{0 \text{ mol}}$   $\underline{\text{xs mol}}$

$$180.0 \text{ g C}_6\text{H}_{12}\text{O}_6 \times \frac{1 \text{ mol C}_6\text{H}_{12}\text{O}_6}{180 \text{ g C}_6\text{H}_{12}\text{O}_6} = 1.00 \text{ mol C}_6\text{H}_{12}\text{O}_6 \times \frac{6 \text{ mol O}_2}{1 \text{ mol C}_6\text{H}_{12}\text{O}_6} = 6.00 \text{ mol O}_2$$

$$6.0 \text{ mol O}_2 \times \frac{32.0 \text{ g O}_2}{1 \text{ mol O}_2} = 192.0 \text{ g O}_2$$

# Answer Key Chemical Equations And Stoichiometry

**John Olmsted, Greg Williams, Robert C.  
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## **Table of Contents Answer Key Chemical Equations And Stoichiometry**

1. Understanding the eBook Answer Key Chemical Equations And Stoichiometry
  - The Rise of Digital Reading Answer Key Chemical Equations And Stoichiometry
  - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key Chemical Equations And Stoichiometry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Answer Key Chemical Equations And Stoichiometry
  - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key Chemical Equations And Stoichiometry
  - Personalized Recommendations
  - Answer Key Chemical Equations And Stoichiometry User Reviews and Ratings
  - Answer Key Chemical Equations And Stoichiometry and Bestseller Lists
5. Accessing Answer Key Chemical Equations And Stoichiometry Free and Paid eBooks
  - Answer Key Chemical Equations And Stoichiometry Public Domain eBooks
  - Answer Key Chemical Equations And Stoichiometry eBook Subscription Services
  - Answer Key Chemical Equations And Stoichiometry Budget-Friendly Options

6. Navigating Answer Key Chemical Equations And Stoichiometry eBook Formats
  - ePub, PDF, MOBI, and More
  - Answer Key Chemical Equations And Stoichiometry Compatibility with Devices
  - Answer Key Chemical Equations And Stoichiometry Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Answer Key Chemical Equations And Stoichiometry
  - Highlighting and Note-Taking Answer Key Chemical Equations And Stoichiometry
  - Interactive Elements Answer Key Chemical Equations And Stoichiometry
8. Staying Engaged with Answer Key Chemical Equations And Stoichiometry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Answer Key Chemical Equations And Stoichiometry
9. Balancing eBooks and Physical Books Answer Key Chemical Equations And Stoichiometry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Answer Key Chemical Equations And Stoichiometry
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Answer Key Chemical Equations And Stoichiometry
  - Setting Reading Goals Answer Key Chemical Equations And Stoichiometry
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Key Chemical Equations And Stoichiometry
  - Fact-Checking eBook Content of Answer Key Chemical Equations And Stoichiometry
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

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