

## Chemistry Worksheets Class 11 on Chapter 4 Chemical Bonding and Molecular Structure - Set 3

**Q-1:** Which of the following possesses an expanded octet?

- a)  $\text{IF}_7$
- b)  $\text{SF}_6$
- c)  $\text{BrF}_5$
- d) All of the above

**Q-2:** Strongly electronegative element B contrasts with strongly electropositive element A. Both are divalent. The resulting compound would be

- a)  $\text{A} \rightarrow \text{B}$
- b)  $\text{A}^{2+}\text{B}^{2-}$
- c)  $\text{A}^{2-}\text{B}^{2+}$
- d)  $\text{A}-\text{B}$

**Q-3:** Which of the following compounds shows the highest lattice energy?

- a)  $\text{CsCl}$
- b)  $\text{KCl}$
- c)  $\text{RbCl}$
- d)  $\text{NaCl}$

**Q-4:** Which one of the following molecules is formed by p-p overlapping?

- a)  $\text{I}_2$
- b)  $\text{HBr}$
- c)  $\text{H}_2\text{S}$
- d)  $\text{PH}_3$

**Q-5:** Which of the following H-bonds has the greatest impact on a molecule's physical characteristics?

- a) Intramolecular H-bond
- b) Intermolecular H-bond
- c) Both of the above
- d) All H-bonds

**Q-6:** State the crucial conditions that must be met for a molecule to undergo hybridisation.

**Q-7:** Describe why  $\text{CH}_4$  has a tetrahedral geometry rather than a square planar geometry with a carbon atom in the centre and four H atoms at each corner.

# Answer Key For Holt Science Chemical Bonding

**Mapi M. Cuevas**



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