

Chemistry, Part 1

Module 14 Notes



Big Question

How are metals used in chemical reactions? and produced by chemical reactions?

Big Idea

Using chemical analysis, you can identify the metals of a substance or pure metal and other substances. A metal is often used in manufacturing. It is used in construction or industrial analysis with chemical reactions.

Learning Objectives

After completing this module, students will be able to:

- Use the law of conservation of mass to balance chemical reactions.
- Use chemical reactions to describe processes of materials.
- Compare different metals and reactions.
- Describe the properties of different metals.

Module 14 Objectives

Learning Objective Use the law of conservation of mass to balance chemical reactions.

Fill in the blanks.

1. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ (Balanced)
2. $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$ (Not balanced)
3. $\text{C}_2\text{H}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ (Not balanced)

Chemical Reactions and Equations

Use the law of conservation of mass to balance chemical reactions.

Chemical reactions are used to produce new materials and products.

Chemical reactions: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$ (Balanced)

Chemical reactions: $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$ (Not balanced)

Chemical reactions: $\text{C}_2\text{H}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ (Not balanced)

Module 14 Objectives

Learning Objective Use the law of conservation of mass to balance chemical reactions.

Fill in the blanks.

1. 1 mole of H_2 reacts to form 1 mole of H_2O .
2. The mass of 1 mole of H_2 is 2 grams. The mass of 1 mole of H_2O is 18 grams.

Answers For Chemistry Module 14

Mr. Rohit Manglik

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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