

Unit 3 Worksheet 3 – Quantitative Energy Problems

Energy constants (H₂O)

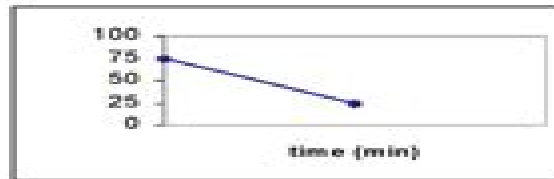
334 J/g	Heat of fusion (melting or freezing) H _f
2260 J/g	Heat of vaporization (evaporating or condensing) H _v
2.1 J/g°C	Heat capacity (c) of solid water
4.18 J/g°C	Heat capacity (c) of liquid water

$$Q = m \cdot c \cdot \Delta T$$

$$Q = m \cdot H_v \text{ or } m \cdot H_f$$

For each of the problems sketch a warming or cooling curve to help you decide which equation(s) to use to solve the problem. Keep a reasonable number of sig figs in your answers.

1. A cup of coffee (140 g) cools from 75°C down to comfortable room temperature 20.°C. How much energy does it release to the surroundings?



$$Q = 140g \cdot 4.18 \frac{J}{g^{\circ}C} \cdot 55^{\circ}C = 32,000J \text{ or } 32 \text{ kJ}$$

2. Suppose during volleyball practice, you lost 2.0 lbs of water due to sweating. If all of this water evaporated, how much energy did the water absorb from your body? Express your answer in kJ. 2.2 lbs = 1.0 kg



$$2.0lbs \cdot \frac{1kg}{2.2lbs} \cdot \frac{1000g}{1kg} = 910g$$

$$Q = 910g \cdot 2260 \frac{J}{g} = 2,100,000J \text{ or } 2100 \text{ kJ}$$

3. Suppose that during the Icy Hot lab that 65 kJ of energy were transferred to 450 g of water at 20°C. What would have been the final temperature of the water?



$$Q = m \cdot c \cdot \Delta T(^{\circ}C)$$

$$\Delta T = \frac{Q}{m \cdot c} = \frac{65,000J}{450g \cdot 4.18 \frac{J}{g^{\circ}C}} = 34.6^{\circ}C = 35^{\circ}C$$

Answer For Quantitative Energy Problems

Puja Mehta



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Table of Contents Answer For Quantitative Energy Problems

1. Understanding the eBook Answer For Quantitative Energy Problems
 - The Rise of Digital Reading Answer For Quantitative Energy Problems
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer For Quantitative Energy Problems
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer For Quantitative Energy Problems
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer For Quantitative Energy Problems
 - Personalized Recommendations
 - Answer For Quantitative Energy Problems User Reviews and Ratings
 - Answer For Quantitative Energy Problems and Bestseller Lists
5. Accessing Answer For Quantitative Energy Problems Free and Paid eBooks
 - Answer For Quantitative Energy Problems Public Domain eBooks
 - Answer For Quantitative Energy Problems eBook Subscription Services
 - Answer For Quantitative Energy Problems Budget-Friendly Options

6. Navigating Answer For Quantitative Energy Problems eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer For Quantitative Energy Problems Compatibility with Devices
 - Answer For Quantitative Energy Problems Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer For Quantitative Energy Problems
 - Highlighting and Note-Taking Answer For Quantitative Energy Problems
 - Interactive Elements Answer For Quantitative Energy Problems
8. Staying Engaged with Answer For Quantitative Energy Problems
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer For Quantitative Energy Problems
9. Balancing eBooks and Physical Books Answer For Quantitative Energy Problems
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer For Quantitative Energy Problems
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer For Quantitative Energy Problems
 - Setting Reading Goals Answer For Quantitative Energy Problems
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer For Quantitative Energy Problems
 - Fact-Checking eBook Content of Answer For Quantitative Energy Problems
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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