

balance nuclear reaction equations

aluminum-26 undergoes positron emission
 potassium-40 undergoes beta decay (electron)
 uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.
 Write the electron configuration for Cr.
 Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride SF_6
- b. Lithium nitride Li_3N
- c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride SnCl_2
- e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- f. Mercury (I) chloride Hg_2Cl_2
- g. Potassium bromate KBrO_3
- h. Hydrosulfuric acid $\text{H}_2\text{S(aq)}$
- i. Chloric acid HClO_3
- j. Sulfurous acid H_2SO_3

Name each of the following compounds:

- a. CuSO_4 Copper (II) sulfate
- b. AlF_3 aluminum fluoride
- c. HI hydrogen iodide
- d. NO nitrogen monoxide
- e. $\text{H}_2\text{Se(aq)}$ hydroselenic acid
- f. HNO_3 nitric acid
- g. HNO_2 nitrous acid
- h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?



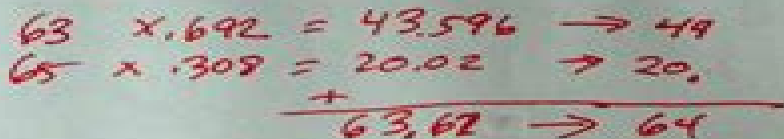
Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?



Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.



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M Lipman



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Table of Contents Answer Key For Study Guide Chemistry

1. Understanding the eBook Answer Key For Study Guide Chemistry
 - The Rise of Digital Reading Answer Key For Study Guide Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key For Study Guide Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key For Study Guide Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key For Study Guide Chemistry
 - Personalized Recommendations
 - Answer Key For Study Guide Chemistry User Reviews and Ratings
 - Answer Key For Study Guide Chemistry and Bestseller Lists

5. Accessing Answer Key For Study Guide Chemistry Free and Paid eBooks
 - Answer Key For Study Guide Chemistry Public Domain eBooks
 - Answer Key For Study Guide Chemistry eBook Subscription Services
 - Answer Key For Study Guide Chemistry Budget-Friendly Options
6. Navigating Answer Key For Study Guide Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Key For Study Guide Chemistry Compatibility with Devices
 - Answer Key For Study Guide Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key For Study Guide Chemistry
 - Highlighting and Note-Taking Answer Key For Study Guide Chemistry
 - Interactive Elements Answer Key For Study Guide Chemistry
8. Staying Engaged with Answer Key For Study Guide Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key For Study Guide Chemistry
9. Balancing eBooks and Physical Books Answer Key For Study Guide Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key For Study Guide Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer Key For Study Guide Chemistry
 - Setting Reading Goals Answer Key For Study Guide Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Key For Study Guide Chemistry
 - Fact-Checking eBook Content of Answer Key For Study Guide Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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