Answers for Predicting Products of Chemical Reactions

For all combustion reactions of hydrocarbons or hydrogen-carbon-oxygen molecules, the products will always be CO: and H:O (assuming a complete reaction).

a. Combustion:	C_6H_{12}	+	901	\rightarrow	6CO ₂		6H;O
b. Combustion:	2C ₄ H ₆	19	1102	+	8CO2	+	6H:O
c. Combustion:	$C_0H_{10}O_3$	94	707		6CO ₂		5H ₂ O

1. Synthesis: Mg + I₂ → MgI:

Note that Mg can only form Mg2 and I can only form I', so the product will must have a 1:2 cation anion ratio.

2. Double displacement: CuCl₂ + H₂S → CuS + 2HCl Note that the product is **not** H₂Cl₂. It is important to recognize that CuCl₂ is made of three ions, Cu¹⁺ and two Cl².

Double displacement: NaOH + HCIO₄ → NaCIO₄ + H₂O
 In this question, you must recognize that perchlorate, ClO₄, and hydroxide, OH¹, are polyatomic ions and will not break apart. Also, this is an acid-base reaction, so the products should be salt and water.

Decomposition: ZnCO₂ + heat → ZnO + CO₂.
 When reactions have heat as a reactant, it is very likely that they will involve decompositions. Carbonate compounds usually decompose to CO₂ and a metal oxide.

5. Single replacement: 2HCl + Zn → ZnCl₂ + H₂
Note that one reactant is in its elemental form; if a single replacement reaction is going to occur, the species NOT in its elemental form in the reactants (H¹ in this case), will end up in its elemental form in the products (H₂) and the species that IS in its elemental form (Zn) will end up ionized (Zn²). Note that zinc can only form a Zn² ion, so it will have two chloride ions. Note also that hydrogen in its elemental form is H₂, not H.

6. <u>Single replacement</u> 2Na + MgCl₂ → 2NaCl + Mg Again notice that one species is in its elemental form (Na). The magnesium in MgCl₂ is an ion (Mg²⁺), but is transformed into its elemental state (Mg), while the Na is converted into an ion (Na⁺, sodium only forms a +1 ion).

Double replacement CaCl: + K:CO → CaCO: + 2KCl
 Recognize that carbonate is a polyatomic ion (CO₁²) and that the cations are already stable ions (Ca²² and K²).

8. <u>Synthesis</u> 2K + CI: → 2KCI

Note that both materials are elemental species, so the only result could be a synthesis reaction. In the product, K^{*} and CF are formed. Note that we would NOT write the product as K₂Cl₂. Ionic compounds are written in their simplest forms.

Double replacement 3BaCl: + 2K.PO₁ → Bas(PO₁): + 6KCl
 Note that phosphate (PO₁) is a polyatomic ion and will not break apart. Since barium is a +2 ion, the barium phosphate will have a 3:2 ratio of Ba:PO₁ in order to balance the charge.

10. <u>Double replacement</u> H:SO_a + 2KOH → 2H:O + K:SO_a Notice that sulfate (SO_a²) is a polyatomic ion and that potassium exists as a +1 ion (K²). Also, this is an acid-base reaction, so the products should be salt and water.

11. Decomposition Al₂(CO₂): + beat → Al₂O₂ + 3CO₂
Another decomposition that will generate CO₂ and a metal oxide (note that Al forms a +3 ion, Al²⁻, and monatomic oxygen will have a -2 charge, O²⁻).

Answer Key Predicting Products Of Chemical Reactions

Ensheng Dong

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these may be supported on bearings although due to durability issues with expansion joints and bearings it is usually preferable to cast them integral with in situ abutments or place them as part of pre cast box culverts