

Honors Chemistry
Extra Stoichiometry Problems

1. Silver nitrate reacts with barium chloride to form silver chloride and barium nitrate.

a. Write and balance the chemical equation.



b. If 39.02 grams of barium chloride are reacted in an excess of silver nitrate, how many representative particles (and what type) of silver chloride are produced?

$$\frac{39.02 \text{ g BaCl}_2}{1} \times \frac{1 \text{ mole BaCl}_2}{208.233 \text{ g BaCl}_2} \times \frac{2 \text{ moles AgCl}}{1 \text{ mole BaCl}_2} \times \frac{6.02 \times 10^{23} \text{ f. u. AgCl}}{1 \text{ mole AgCl}} = 2.256 \times 10^{23} \text{ f. u. AgCl}$$

c. If 410.8 grams of barium nitrate are produced how many grams of silver nitrate were reacted?

$$\frac{410.8 \text{ g Ba}(\text{NO}_3)_2}{1} \times \frac{1 \text{ mole Ba}(\text{NO}_3)_2}{261.335 \text{ g Ba}(\text{NO}_3)_2} \times \frac{2 \text{ moles AgNO}_3}{1 \text{ mole Ba}(\text{NO}_3)_2} \times \frac{169.872 \text{ g AgNO}_3}{1 \text{ mole AgNO}_3} = 534.1 \text{ g AgNO}_3$$

2. Nitrogen gas is reacted with hydrogen gas to form nitrogen trihydride.

a. Write and balance the chemical equation.



b. How many liters of nitrogen trihydride are produced at STP if 80.28 grams of hydrogen gas are reacted in an excess of nitrogen?

$$\frac{80.28 \text{ g H}_2}{1} \times \frac{1 \text{ mole H}_2}{2.016 \text{ g H}_2} \times \frac{2 \text{ moles NH}_3}{3 \text{ moles H}_2} \times \frac{22.4 \text{ L NH}_3}{1 \text{ mole NH}_3} = 594.7 \text{ L NH}_3$$

c. How many grams of hydrogen are needed to fully react 621.9 grams of nitrogen gas?

$$\frac{621.9 \text{ g N}_2}{1} \times \frac{1 \text{ mole N}_2}{28.014 \text{ g N}_2} \times \frac{3 \text{ moles H}_2}{1 \text{ mole N}_2} \times \frac{2.016 \text{ g H}_2}{1 \text{ mole H}_2} = 134.3 \text{ g H}_2$$

d. How many liters of nitrogen gas are needed to completely react 90.38 L of hydrogen gas at STP?

$$\frac{90.38 \text{ L H}_2}{1} \times \frac{1 \text{ mole H}_2}{22.4 \text{ L H}_2} \times \frac{1 \text{ mole N}_2}{3 \text{ mole H}_2} \times \frac{22.4 \text{ L N}_2}{1 \text{ mole N}_2} = 30.09 \text{ L N}_2$$

$$\begin{array}{c} \text{OR} \\ \frac{90.38 \text{ L H}_2}{1} \times \frac{1 \text{ L N}_2}{3 \text{ L H}_2} = 30.09 \text{ L N}_2 \end{array}$$

Answer Key Stoichiometry Practice Problems

S Ben Porath



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Table of Contents Answer Key Stoichiometry Practice Problems

1. Understanding the eBook Answer Key Stoichiometry Practice Problems
 - The Rise of Digital Reading Answer Key Stoichiometry Practice Problems
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key Stoichiometry Practice Problems
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key Stoichiometry Practice Problems
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key Stoichiometry Practice Problems
 - Personalized Recommendations
 - Answer Key Stoichiometry Practice Problems User Reviews and Ratings
 - Answer Key Stoichiometry Practice Problems and Bestseller Lists

5. Accessing Answer Key Stoichiometry Practice Problems Free and Paid eBooks
 - Answer Key Stoichiometry Practice Problems Public Domain eBooks
 - Answer Key Stoichiometry Practice Problems eBook Subscription Services
 - Answer Key Stoichiometry Practice Problems Budget-Friendly Options
6. Navigating Answer Key Stoichiometry Practice Problems eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Key Stoichiometry Practice Problems Compatibility with Devices
 - Answer Key Stoichiometry Practice Problems Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Answer Key Stoichiometry Practice Problems
 - Highlighting and Note-Taking Answer Key Stoichiometry Practice Problems
 - Interactive Elements Answer Key Stoichiometry Practice Problems
8. Staying Engaged with Answer Key Stoichiometry Practice Problems
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key Stoichiometry Practice Problems
9. Balancing eBooks and Physical Books Answer Key Stoichiometry Practice Problems
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key Stoichiometry Practice Problems
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer Key Stoichiometry Practice Problems
 - Setting Reading Goals Answer Key Stoichiometry Practice Problems
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Key Stoichiometry Practice Problems
 - Fact-Checking eBook Content of Answer Key Stoichiometry Practice Problems
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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