

Electrostatic Potential Lab

Name _____ Per _____

Circle the molecule below: Circle the ring carbon in the structure.

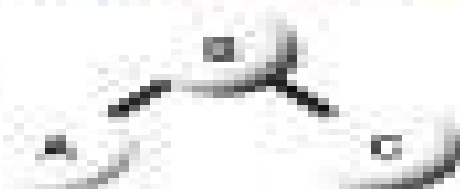
Part A: Three atom investigation

1. Click the "Keep the box" button adjacent to the three atoms and click the "Electrostatic Potential" button.
2. Change the electronegativity of **A** and **B** to the electronegativity of atom **B**. Observe the color, partial charges and bond structure. Write the following information: the electronegativity of atom B (use the top of the page) _____
 partial charges _____
 bond structure _____

3. Now put the electronegativity of A and B into the electronegativity of B and observe the results. Fill in the following information:
 - a. The polarity arrow points to the _____ electronegative atom.
 - b. The partial positive charge is always on the _____ electronegative atom.
 - c. The larger the electronegativity difference, the _____ the bond structure.

Part B: Three atom investigation

1. Change the electronegativity of atom A and B to the electronegativity of **A** and **C**. Observe the color, partial charges, molecular dipole and partial charges.
 - a. Draw the partial charges on the Lewis dot diagram.
 - b. In words and color, describe what molecular dipole is in the diagram.
 - c. Draw the partial charge symbols and position them in the diagram.
 - d. Circle the electronegative. Add other electronegative the molecule, if appropriate.
2. Click the "Keep the box" button. Change the electronegativity of **A** and **B** to the electronegativity of **B** and **C**. Observe the color, partial charges and partial charges.
 - a. Draw the partial charges on the Lewis dot diagram.
 - b. In words and color describe what molecular dipole is in the diagram.
 - c. Draw the partial charge symbols and position them in the diagram.
 - d. Circle the electronegative. Add other electronegative the molecule, if appropriate.



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