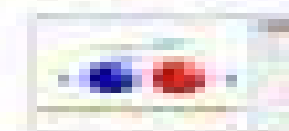
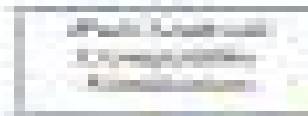


Experiment 1: The Effect of Temperature on the Rate of Reaction

Name: _____

Background Information

Every science experiment is designed to test a hypothesis. In this experiment, the hypothesis is that the rate of reaction will increase as the temperature of the reactants increases. The rate of reaction is the change in concentration of the products per unit time. The rate of reaction can be measured by the change in the amount of the products formed over a given time interval. The rate of reaction can also be measured by the change in the amount of the reactants consumed over a given time interval. The rate of reaction is affected by many factors, including temperature, concentration, and surface area. In this experiment, the effect of temperature on the rate of reaction will be studied.



Procedure

1. Prepare a solution of 0.1 M H_2O_2 and 0.1 M KI in a beaker.

2. Measure the volume of the solution.

3. Add a small amount of MnO_2 to the solution.

Results

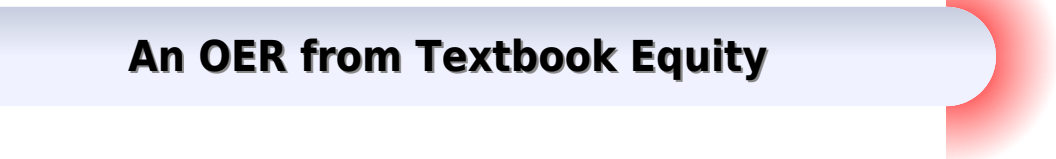
Table 1.1

1. The rate of reaction increases as the temperature increases.
2. The rate of reaction increases as the concentration of the reactants increases.
3. The rate of reaction increases as the surface area of the reactants increases.

Time (s)	Volume of O_2 (mL)	Rate of Reaction (mL/s)	Temperature (°C)	Concentration of H_2O_2 (M)
0	0	0	20	0.1
10	10	1.0	20	0.1
20	20	2.0	20	0.1
30	30	3.0	20	0.1
40	40	4.0	20	0.1
50	50	5.0	20	0.1
60	60	6.0	20	0.1
70	70	7.0	20	0.1
80	80	8.0	20	0.1
90	90	9.0	20	0.1
100	100	10.0	20	0.1

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