

3. Ionic bond is more solid where covalent bond has a lower melting point , and ionic bonds result from transfer of electrons , whereas covalent bonds are formed by sharing

7. the bonding atoms obtain stable electron configurations; since the stable configuration is usually eight electrons in the outermost shell

15. A shared pair of electrons is called a **bonding pair**, while a pair that is associated with only one atom—and therefore not involved in bonding—is called a **lone pair**.

24. Atoms and molecules , CO₂

29.

- a. molecular
- b. ionic
- c. ionic
- d. molecular

33.

- a. 3 Mg , 2 P , 8 O
- b. 1 Ba , 2 Cl
- c. 1 Fe , 2 N , 4 O
- d. 1 Ca , 2 O , 2 H

39.



43.

- a. SrSe
- b. BaCl₂
- c. Na₂S
- d. Al₂O₃

45. The product of the charges of the two ions is one factor of lattice energy. CsF has a +1 ion charge product, whereas BaO has a +4 ion charge product. The lattice energy is four times bigger because the product is four times greater.

57.

- a. NaHSO₃
- b. LiMnO₄
- c. AgNO₃
- d. K₂SO₄

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Matthew E. Johll



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