

Key: Half-Life Calculations Practice

First fill in the blank with the correct isotope symbol to complete the nuclear equations. Then apply what you know about half-life to answer the questions.

Half-life = $t_{1/2}$ = the time it takes for half of the original sample to decay into another element.

1. What mass of a 100. g sample remains as U-238 after 60. Sec. $T_{1/2} = 20$ sec.
 $^{238}_{92}\text{U} \rightarrow ^{234}_{90}\text{Th} + ^{4}_{2}\text{He}$
25.00 g
2. Of an original mass of 100. g, how much remains as Pb-214 after 48,210 yrs.?
 $^{214}_{82}\text{Pb} \rightarrow ^{214}_{81}\text{Bi} + ^{0}_{-1}\text{e}^-$ $T_{1/2} = 26,810$ yrs
12.5 g
3. What fraction of a given amount of protactinium 231 remains after 26.76 hrs.?
 $^{231}_{91}\text{Pa} \rightarrow ^{231}_{90}\text{Th} + ^{0}_{+1}\text{e}^+$ $T_{1/2} = 6.69$ hrs
1/16 of the sample remains
4. How many milligrams of a 15.0 mg sample of radium-226 remains after 3198 yrs.?
 $^{226}_{88}\text{Ra} \rightarrow ^{226}_{87}\text{Fr} + ^{4}_{2}\text{He}$ $T_{1/2} = 1599$ yrs
3.75 mg
5. How much of an original 0.250 g sample of radium-226 remains after 1599 yrs.?
 $^{226}_{88}\text{Ra} \rightarrow ^{226}_{87}\text{Fr} + ^{4}_{2}\text{He}$ $T_{1/2} = 1599$ yrs
0.0613 g
6. What was the original mass of radium-224 if 0.0500 g remains after 7.32 days?
 $^{224}_{88}\text{Ra} \rightarrow ^{224}_{87}\text{Fr} + ^{4}_{2}\text{He}$ $T_{1/2} = 3.66$ days
0.200 g

12 Tesccc Chemistry Half Life Practice Answers

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