

For questions 1 - 7, choose true or false.

1. The rate of a given reaction is constant at any specified temperature. **False**
2. A balanced equation does not indicate the mechanism by which a reaction is proceeding. **True**
3. The rate constant is the rate of reaction at constant temperature. **False**
4. In the reaction $\text{H}_2 + \text{I}_2 \rightleftharpoons 2 \text{HI}$, only two molecules (H_2 and I_2) react so this reaction is a bimolecular reaction. **False**
5. Relatively few reaction mechanisms have been completely characterized. **True**
(Special ingenuity and techniques are needed to detect fast steps in a reaction mechanism.)
6. The rate-determining step of a reaction is the step that proceeds fastest. **False**
7. The half-life of a reaction is the time required for half of the original concentration of the limiting reagent to be consumed in the reaction. **True**
8. Factors that determine reaction rates include all of the following except:
a. size of reactant particles
b. **magnitude of the equilibrium constant** (Large values of k do not imply large rates.)
c. reaction temperature
d. concentration of reactants
9. For the reaction $2\text{A} + \text{B} \rightarrow \text{C}$, the reaction rate is:
a. $R = k[\text{A}]^2 [\text{B}]$
b. $R = k[\text{A}] [\text{B}]$
c. $R = k[\text{A}]^2$
d. **possibly not any of the above** (A valid reaction rate expression usually cannot be written from the overall reaction.)
10. For a reaction of the type $\text{A} + \text{B} \rightarrow 2\text{C}$, it is found that doubling the amount of A causes the reaction rate to be four times as great but that doubling the amount of B has no apparent effect on the rate. The rate equation that is most correct is:
a. **$R = k[\text{A}]^2$ (second order in A and zero order in B)**
b. $R = k[\text{A}] [\text{B}]$
c. $R = k[\text{A}]$
d. $R = \frac{k[\text{C}]^2}{[\text{A}] [\text{B}]}$
11. Which of the following is the rate equation for a termolecular elementary reaction?
a. $R = k [\text{A}]$
b. $R = k [\text{A}]^2$
c. $R = k [\text{A}] [\text{B}]$
d. **$R = k [\text{A}] [\text{B}] [\text{C}]$ (three component collision)**

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