

The mixture of the hydrocarbon and oxygen is sparked so that a complete combustion reaction occurs, producing $\text{CO}_2(g)$ and $\text{H}_2\text{O}(g)$. The partial pressures of these gases at 127°C are 0.600 atm for $\text{CO}_2(g)$ and 0.800 atm for $\text{H}_2\text{O}(g)$. There is $\text{O}_2(g)$ remaining in the container after the reaction is complete.

(c) Use the partial pressures of $\text{CO}_2(g)$ and $\text{H}_2\text{O}(g)$ to calculate the partial pressure of the $\text{O}_2(g)$ consumed in the combustion.

(d) On the basis of your answers above, write the balanced chemical equation for the combustion reaction and determine the formula of the hydrocarbon.

c)
$$\text{---C}_x\text{H}_y + \text{---}\underline{\text{O}_2} \rightarrow \text{---}\underline{\text{CO}_2} + \text{---}\underline{\text{H}_2\text{O}}$$

$$0.6\text{ atm CO}_2 \cdot \frac{1\text{ O}_2}{2\text{ CO}_2} = 0.3\text{ atm O}_2$$

$$0.8\text{ atm H}_2\text{O} \cdot \frac{1\text{ O}_2}{2\text{ H}_2\text{O}} = 0.4\text{ atm O}_2$$

$$\left. \begin{array}{l} 0.3\text{ atm O}_2 \\ 0.4\text{ atm O}_2 \end{array} \right\} 0.7\text{ atm O}_2$$

Extra Page

2012 Ap Chemistry Response Answer Key

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