

Name: _____
Date: _____ Period: _____

Chemistry: *The Periodic Table*

Directions: Answer each of the following questions. You need not use complete sentences.

What is a period? How many are there in the periodic table?

A period is a row of elements on the periodic table, in which the elements are ordered one proton more each place from left to right. A period also often helps to represent an electron energy level and the order in which electrons are filled into their sublevels. There are 7 periods.

What is a group (also called a family)? How many are there in the periodic table?

Groups are the columns that are found top to bottom in the table. They show elements that have the same sublevel and number of electrons in their outer shell, valence electrons. Often elements in a group share any similar chemical properties because of this. There are 18 groups in the periodic table.

State the number of valence electrons in an atom of:

a. sulfur
6

b. calcium
2

c. chlorine
7

d. arsenic
5

Give the names and chemical symbols for the elements that correspond to these atomic numbers:

a. 10

Neon, Ne

b. 18

Argon, Ar

c. 36

Krypton, Kr

d. 90

Thorium, Th

List, by number, both the period and group of each of these elements.

	Symbol	Period	Group
a. beryllium	Be	2	2
b. iron	Fe	4	8
c. lead	Pb	6	14

Which of the following pairs of elements belong to the same period?

a. Na and Cl

b. Na and Li

c. Na and Cu

d. Na and Ne

Which of the following pairs of elements belong to the same group?

a. H and He

b. Li and Be

c. C and Pb

d. Ga and Ge

Would you expect strontium to be, chemically, more similar to calcium or rubidium and WHY?

Calcium, because they share a common amount of valence electrons. Valence electrons are the main contributor to chemical properties and how elements interact with other matter. Both strontium and calcium have two electrons on the the "s" sublevel, $5s^2$ and $4s^2$

What are the Group 1 elements called?

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J Ma



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Table of Contents Answers Key To Chemistry Radioactivity Packet 15

1. Understanding the eBook Answers Key To Chemistry Radioactivity Packet 15
 - The Rise of Digital Reading Answers Key To Chemistry Radioactivity Packet 15
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers Key To Chemistry Radioactivity Packet 15
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers Key To Chemistry Radioactivity Packet 15
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers Key To Chemistry Radioactivity Packet 15
 - Personalized Recommendations
 - Answers Key To Chemistry Radioactivity Packet 15 User Reviews and Ratings
 - Answers Key To Chemistry Radioactivity Packet 15 and Bestseller Lists
5. Accessing Answers Key To Chemistry Radioactivity Packet 15 Free and Paid eBooks

- Answers Key To Chemistry Radioactivity Packet 15 Public Domain eBooks
- Answers Key To Chemistry Radioactivity Packet 15 eBook Subscription Services
- Answers Key To Chemistry Radioactivity Packet 15 Budget-Friendly Options
- 6. Navigating Answers Key To Chemistry Radioactivity Packet 15 eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers Key To Chemistry Radioactivity Packet 15 Compatibility with Devices
 - Answers Key To Chemistry Radioactivity Packet 15 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers Key To Chemistry Radioactivity Packet 15
 - Highlighting and Note-Taking Answers Key To Chemistry Radioactivity Packet 15
 - Interactive Elements Answers Key To Chemistry Radioactivity Packet 15
- 8. Staying Engaged with Answers Key To Chemistry Radioactivity Packet 15
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers Key To Chemistry Radioactivity Packet 15
- 9. Balancing eBooks and Physical Books Answers Key To Chemistry Radioactivity Packet 15
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers Key To Chemistry Radioactivity Packet 15
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers Key To Chemistry Radioactivity Packet 15
 - Setting Reading Goals Answers Key To Chemistry Radioactivity Packet 15
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers Key To Chemistry Radioactivity Packet 15
 - Fact-Checking eBook Content of Answers Key To Chemistry Radioactivity Packet 15
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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