

Homework 13.1| Covalent Bond Name:**Grade and Section:****Directions:** Complete the table with the information needed. Refer to the periodic table below if needed.**Table A**

Element	Group no.	No. Electrons	No. of Valence Electrons	Lewis Dot Symbol
O	6A	8	6	$\text{:}\ddot{\text{O}}\cdot$
I	7A	53	7	$\text{:}\ddot{\text{I}}\cdot$
Cl	7A	17	7	$\text{:}\ddot{\text{Cl}}\cdot$
N	5A	7	5	$\text{:}\ddot{\text{N}}\cdot$
P	5A	15	5	$\text{:}\ddot{\text{P}}\cdot$
C	4A	6	4	$\text{:}\ddot{\text{C}}\cdot$
S	6A	16	6	$\text{:}\ddot{\text{S}}\cdot$
Br	7A	35	7	$\text{:}\ddot{\text{Br}}\cdot$
F	7A	9	7	$\text{:}\ddot{\text{F}}\cdot$

Table B

Compound	Lewis structure	Bond Polarity	Polarity of Molecules
1. CO_2	$\text{:}\ddot{\text{O}}=\text{C}=\ddot{\text{O}}\text{:}$	$\text{O}-\text{C } 3.5 - 2.5 = 1$	Polar
2. CH_4	$ \begin{array}{c} \text{H} \\ \\ \text{H}-\text{C}-\text{H} \\ \\ \text{H} \end{array} $	$\text{C}-\text{H } 2.5 - 2.1 = 0.4$	Non-Polar
3. PCl_5	$ \begin{array}{c} \text{:}\ddot{\text{Cl}}\text{:} \\ \\ \text{:}\ddot{\text{Cl}}-\text{P}-\ddot{\text{Cl}}\text{:} \\ / \quad \backslash \\ \text{:}\ddot{\text{Cl}}\text{:} \quad \text{:}\ddot{\text{Cl}}\text{:} \end{array} $	$\text{Cl}-\text{P } 3.0 - 2.1 = 0.9$	Polar

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