Stoichiometry Worksheet #2 (mole-mass, mass-mole problems)

- N₂ + 2O₂ → N₂O₄
 - a. If 15.0g of N₂O₄ was produced, how many moles of O₂ were required?

15.0g N₂O₄ 1 mol N₂O₄ 2 mol O₂ = 0.326 mol O₂ 92.0g N₂O₄ 1 mol N₂O₄ = 0.326 mol O₂

b. If 4.0x10⁻³ moles of oxygen reacted, how many grams of N₂ were needed?

- K₃PO₄ + AJ(NO₃)₃ → 3KNO₃ + AJPO₄
 - a. What is the mass of potassium nitrate that is produced when 2.04 moles of potassium phosphate react?

2.04 mol K₃PO₄ 3 mol KNO₃ 101.1 g KNO₃ = 619g KNO₃ 1 mol K₃PO₄ 1 mol KNO₃

> b. If 5.80g of aluminum phosphate are formed, how many moles of aluminum nitrate reacted?

5.80g AIPO₄ 1 mol AIPO₆ 1 mol AI(NO₁)₁ = 4.75×10⁻² mol AI(NO₁)₁ 122.0g AIPO₆ 1 mol AIPO₆

- CaC₂ + 2H₂O → Ca(OH)₂ + C₂H₂
 - a. If you have 5.50 mol of CaC₂, how much C₂H₂ do you get?

5.50 mol CaC₂ 1 mol C₂H₂ 26.0g C₂H₂ = 143g C₂H₂ 1 mol CaC₂ 1 mol C₂H₂

b. How many moles of water are needed when 65.0g of CaC₂ have reacted?

65.0g CaC₂ 1 mol CaC₂ 2 mol H₂O = 2.03 mol H₂O 64.1g CaC₂ 1 mol CaC₂

4. In photosynthesis, water reacts with carbon dioxide to give oxygen and glucose (C₆H₁₂O₆). Write and balance the chemical equation. How many moles of CO₂ are required to make 120.0g of glucose?

6H2O + 6CO2 -- C4H22O4 + 6O2

120.0g C₆H₁₂O₈ 1 mol C₆H₁₂O₈ 6 mol CO₂ = 4.000 mol CO₂ 180.0 g C₆H₁₂O₈ 1 mol C₆H₁₂O₆

Answers For The Chemistry Lab Of Stoichiometry

Joseph A. Mascetta, Mark Kernion

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the experimental nature of general chemistry Experiments are written to address students of various academic backgrounds and differing interests and abilities in chemistry While most experiments can be conducted in a single three hour period some have been designed to be completed over an extended time to illustrate that chemical systems do not work at an arbitrary schedule Suggestions are provided for combining experiments of shorter length and similar pedagogy

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Energy Research Abstracts ,1989

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