

SCH4U Unit 4 - Quiz – Equilibrium

Name: _____ Period: _____

Q1: (10)

Q2: (10)

Q3: (10)

Q4: (10)

Q5: (10)

Write the formulas for:

$$K_{eq} = \frac{[products]}{[reactants]} \quad ; \quad K_{eq} = \frac{P_{products}}{P_{reactants}}$$

$$K_{eq} = \frac{P_{CO_2}^2 P_{H_2O}}{P_{CH_4} P_{H_2O}^2} \quad ; \quad K_{eq} = \frac{P_{CO_2}^2 P_{H_2O}}{(P_{CH_4}) (P_{H_2O})^2}$$

gpt (aq) + h₂o (l) ⇌ gpt(s) + h₂g (g) ; gpt(s) ⇌ h₂g (g)

K_{eq} = 1.0 x 10⁻⁴

Q6: (10)

6. Hydrogen sulfide, H₂S, is a slightly poisonous gas. At 100°C, an equilibrium mixture contains 0.020 moles of hydrogen, H₂, 0.010 moles of hydrogen sulfide, and an undetermined amount of sulfur, S₂. The reaction is:



If the volume of the mixture is 1.0 L, what amount of S₂ is present at equilibrium?

a. 0.005

b. 0.001 moles of S₂

c. 0.001 moles of S₂

d. 0.001 moles of S₂

e. Determine the amount of S₂ at equilibrium

f. 0.001

What other species have chemical formula consisting of 2 atoms containing sulfur?

g. 0.001

	2H ₂ (g)	S ₂ (g)	2H ₂ S(g)
Initial (moles)			
Change (moles)			
Final (moles)	0.020	0.010	

f. 0.001 moles of S₂

Answers For Sch4u Unit 4

Tsfx



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