AP® CHEMISTRY 2004 SCORING GUIDELINES

Question 1

Answer the following questions relating to the solubilities of two silver compounds, Ag₂CrO₄ and Ag₃PO₄.

Silver chromate dissociates in water according to the equation shown below.

$$Ag_2CrO_4(s) \rightleftharpoons 2 Ag^*(aq) + CrO_4^{2-}(aq)$$
 $K_{sp} = 2.6 \times 10^{-12} \text{ at } 25^{\circ}C$

(a) Write the equilibrium-constant expression for the dissolving of Ag₂CrO₄(s).

$K_{sp} = [Ag^{+}]^{2}[CrO_{4}^{2}]$	1 point for correct expression	
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(b) Calculate the concentration, in mol L⁻¹, of Ag⁺(aq) in a saturated solution of Ag₂CrO₄ at 25°C.

$Ag^{\dagger}(aq)$
ng [Ag*]
n

(c) Calculate the maximum mass, in grams, of Ag₂CrO₄ that can dissolve in 100. mL of water at 25°C.

$\frac{8.7 \times 10^{-5} \text{ mole Ag}_2\text{CrO}_4}{1 \text{ L}} \times \frac{331.7 \text{ g}}{1 \text{ mole Ag}_2\text{CrO}_4} = \frac{0.029 \text{ g Ag}_2\text{CrO}_4}{1 \text{ L}}$	1 point for molar mass of Ag ₂ CrO ₄
$\frac{0.029 \text{ g Ag}_2\text{CrO}_4}{1 \text{ L}} \times 0.100 \text{ L} = 0.0029 \text{ g Ag}_2\text{CrO}_4$	1 point for mass of Ag ₂ CrO ₄ in 100 mL

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