

Chemistry, Part 1

Module 14 Notes



Big Question

How are metals used in chemical reactions? and produced by chemical reactions?

Big Idea

Using chemical analysis, you can identify the metals of a substance or pure metal and other substances. A metal is often used in manufacturing. It is used in construction or industrial analysis with chemical reactions.

Learning Objectives

After completing this module, students will be able to:

- Use the law of conservation of mass to balance chemical reactions.
- Use chemical reactions to describe processes of materials.
- Compare different metals and reactions.
- Describe the properties of different metals.

Module 14 Objectives

Learning Objective Use the law of conservation of mass to balance chemical reactions.

Fill in the blanks.

1. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ (Balanced)
2. $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$ (Balanced)
3. $\text{K}_2\text{SO}_4 + \text{BaCl}_2 \rightarrow \text{BaSO}_4 + 2\text{KCl}$ (Balanced)

Chemical Reactions and Equations

Use the law of conservation of mass to balance chemical reactions.

Chemical reactions are the chemical processes that produce new substances.

Chemical reactions: $\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$ (Balanced)

Chemical reactions: $\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$ (Balanced)

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Module 14 Objectives

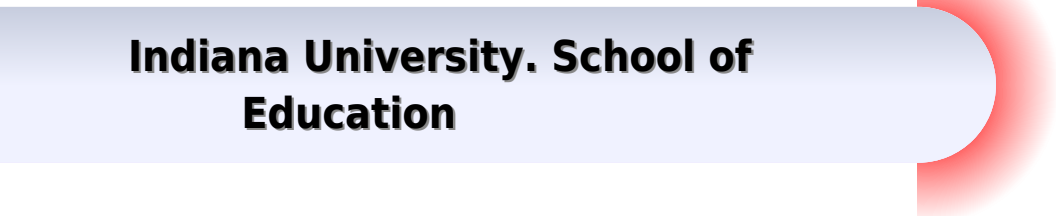
Learning Objective Use the law of conservation of mass to balance chemical reactions.

Fill in the blanks.

1. 1 mole of H_2 reacts to form 1 mole of H_2O .
2. The mass of 1 mole of H_2 is 2.016 g. The mass of 1 mole of H_2O is 18.015 g.

Answers For Chemistry Module 14

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