

1. Answer: A

Explanation: The value of ΔH is negative. Therefore, the reaction is exothermic or heat-releasing.

2. Answer: A

Explanation: The entire reaction was multiplied by $\frac{1}{2}$. Therefore, ΔH must also be halved.

3. Answer: C

Explanation: Option C is incorrect because -566 kJ/mol is the ΔH per mole of O_2 and not for 2 moles. Even though options B and D are expressed in grams, they are still correct because 56 g of CO still corresponds to 2 moles of CO, and 32 g of O_2 still corresponds to 1 mole of O_2 , which is consistent with the expected factor.

4. Answer: A

Explanation: $\Delta H = (98 \text{ g CO}_{(g)}) \left(\frac{1 \text{ mol CO}_{(g)}}{28 \text{ g CO}_{(g)}} \right) \left(\frac{-566 \text{ kJ}}{1 \text{ mol CO}_{(g)}} \right) = -990.5 \text{ kJ}$

5. Answer: B

Explanation: $\Delta H = (1.5 \text{ mol CO}_{2(g)}) \left(\frac{-566 \text{ kJ}}{1 \text{ mol CO}_{2(g)}} \right) = -424.5 \text{ kJ}$

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H Kauffman



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