-Grade 11 chemistry pg. 41 Q1-3 5-6

1. identify which element in each of the following groups has the largest atomic radius:

alikali metals

2. Arrange the following atoms in order of increasing atomic radius (smallest to largest): K, Rb, Cs, F, Li, C

F>C>Li>K>Rb>Cs

3. Predict which has the larger radius: a magnesium atom or a magnesium ion. Explain.

Magnesium atom would have the larger radius that's because a magnesium ion has more charges so it attracts the nucleus towards it self therefore making it smaller.

 For each of the following pairs of atoms, state which would have (a) a larger size (b) a higher ionization energy (c) a lower electron affinity
Ca, K (ii) AI, P (iii) I, F (iv) Li, Ra

(i) Ca,k:

a) k

b) k

c) car

(iii) Al, P:

a) p

6) p

c) all

(100) I,F:

a) f

PIL

c) 1

(IV) LL,Rac

a) ra

b) ra

C) B

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