

SCHAU Unit 4 - Quiz – Equilibrium

Player(s): _____

Period(s): _____

Q1: (10)

Q2: (10)

Q3: (10)

Q4: (10)

Q5: (10)

Useful formulae:

$$K_c = \frac{[\text{products}]}{[\text{reactants}]} \quad ; \quad K_p = \frac{p_{\text{products}}}{p_{\text{reactants}}}$$

$$K_c = \frac{p_{\text{H}_2}^2 p_{\text{O}_2}}{p_{\text{H}_2\text{O}}^2} \quad ; \quad K_p = \frac{p_{\text{H}_2}^2 p_{\text{O}_2}}{(p_{\text{H}_2\text{O}})^2}$$

$$p_{\text{H}_2} = p_{\text{H}_2\text{O}} \times K_c \quad ; \quad p_{\text{O}_2} = K_c \times p_{\text{H}_2\text{O}} \quad ; \quad p_{\text{H}_2\text{O}} = K_p \times p_{\text{H}_2}^2 \times p_{\text{O}_2}$$

$$K_p = K_c \times (RT)^{\Delta n}$$

Question 1:

1. Hydrogen sulfide, H_2S , is a slightly poisonous gas. At 1000 K, an equilibrium mixture contains 0.020 moles of hydrogen, H_2 , 0.010 moles of hydrogen sulfide, with an equilibrium pressure of 0.600 atm. The reaction is:



If the volume of an 8.00 L flask at 1000 K where a concentration of H_2 is present at equilibrium:

- What is the concentration of H_2 at equilibrium?
 Answer: $[\text{H}_2] = 0.0025 \text{ mol/L}$
 Answer: $[\text{H}_2] = 0.020 \text{ mol/L}$
 Answer: $[\text{H}_2] = 0.010 \text{ mol/L}$
- Calculate the value of K_c at equilibrium.
- Answer:
 What other species have chemical formulae associated with this equilibrium?
- Calculate:

	$\text{H}_2\text{S}_{(g)}$	$\text{H}_{2(g)}$	$\text{S}_{2(g)}$
Initial concentration (mol/L)			
Change in concentration (mol/L)			
Equilibrium concentration	0.020	0.010	0.005

What is the value of K_c ?

Answers For Sch4u Unit 4

Rachel S Tattersall



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