

QUESTION 2 (contd)

(b. iii) The system is an exothermic one, when it is cooled, the equilibrium will adjust to annul the cooling effect by releasing more heat. And this favours the forward reaction.

$$\text{(c) i) } \text{PbCl}_2 \rightarrow \text{Pb}^{2+} + 2\text{Cl}^-$$

$$\text{No of moles of PbCl}_2 = \frac{\text{mass}}{\text{molar mass}} = \frac{92.7}{207 + (35.5 \times 2)}$$

$$= \frac{92.7}{278}$$

$$= 0.33345 \text{ moles}$$

$$\text{No of moles of Pb}^{2+} \text{ ionized} = 0.33345 \text{ moles}$$

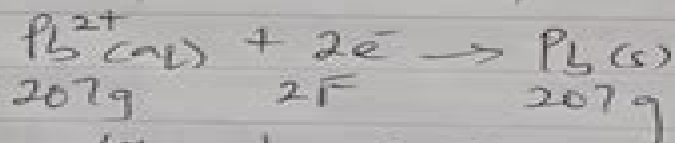
$$\text{mass of lead present} = \text{no of moles} \times \text{molar mass}$$

$$= 0.33345 \times 207$$

$$= 69.0 \text{ grams}$$

$$\text{(ii) Quantity of electricity used} = It = I \times 4.8 \times 3600$$

$$= (17280 \text{ I}) \text{ C}$$



2×96500 liberate 207 grams of Pb

17280I will liberate $\frac{17280I \times 207}{2 \times 96500} = 207$

$$I = \frac{2 \times 96500 \times 207}{17280 \times 207} = 11.17 \text{ Amperes}$$

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