

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive an empirical formula from a chemical structure without performing stoichiometric calculations.
2. To calculate the number of moles of a solid from a weighing measurement in the lab.
3. To determine the molar mass of a solid.
4. To determine the theoretical amount of product that would be produced.
5. To compare the theoretical amount to the actual amount of substance formed and calculate the percent yield.

Reaction equation:

Reaction of sodium chloride solution: 25.0 mL

Reaction of calcium chloride solution: 25.0 mL

Reaction of 10% NaOH: 5.00 g

Reaction of 10% NaOH: 5.00 g

Reaction of 10% NaOH: 5.00 g

Reaction of 10% NaOH: 5.00 g

Procedure:

Reaction and Calculations:

1. Reaction:

$$\text{mole of sodium chloride} = 25.0 \text{ mL} \times \frac{1.0 \text{ g/mL}}{58.44 \text{ g/mol}} = 0.428 \text{ mol}$$

$$\text{mole of calcium chloride} = 25.0 \text{ mL} \times \frac{1.0 \text{ g/mL}}{110.98 \text{ g/mol}} = 0.225 \text{ mol}$$

$$\text{mole of sodium chloride produced} = 0.428 \text{ mol} \times \frac{1 \text{ mol}}{1 \text{ mol}} = 0.428 \text{ mol}$$

$$\text{mole of calcium chloride produced} = 0.225 \text{ mol} \times \frac{1 \text{ mol}}{1 \text{ mol}} = 0.225 \text{ mol}$$

2. Reaction:

Reaction:

$$\text{mole of } \text{CaCl}_2 = 0.225 \text{ mol} \times \frac{1 \text{ mol}}{1 \text{ mol}} = 0.225 \text{ mol}$$

$$\text{mole of } \text{CaCl}_2 = 0.225 \text{ mol} \times \frac{1 \text{ mol}}{1 \text{ mol}} = 0.225 \text{ mol}$$

$$\text{mole of } \text{CaCl}_2 = 0.225 \text{ mol} \times \frac{1 \text{ mol}}{1 \text{ mol}} = 0.225 \text{ mol}$$

Answers For Lab 6b Heath Chemistry

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S. Coelho, Bin Tan, Robert Franzén**

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