



Redox Reactions

Name _____

Date _____

Balance the following redox reactions.

1. $\text{H}_2\text{SO}_3 + \text{I}_2 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{SO}_4 + \text{HI}$
2. $\text{FeCl}_3 + \text{SnCl}_2 \rightarrow \text{FeCl}_2 + \text{SnCl}_4$
3. $\text{SbCl}_5 + \text{KI} \rightarrow \text{SbCl}_3 + \text{KCl} + \text{I}_2$
4. $\text{CdS} + \text{I}_2 + \text{HCl} \rightarrow \text{CdCl}_2 + \text{HI} + \text{S}$
5. $\text{KClO}_3 + \text{FeSO}_4 + \text{H}_2\text{SO}_4 \rightarrow \text{KCl} + \text{Fe}_2(\text{SO}_4)_3 + \text{H}_2\text{O}$
6. $\text{CuSO}_4 + \text{KI} \rightarrow \text{CuI} + \text{K}_2\text{SO}_4 + \text{I}_2$
7. $\text{AuCl}_3 + \text{KI} \rightarrow \text{AuCl} + \text{KCl} + \text{I}_2$
8. $\text{SnCl}_2 + \text{HgCl}_2 \rightarrow \text{SnCl}_4 + \text{HgCl}$
9. $\text{Ti}_2(\text{SO}_4)_3 + \text{FeI}_2(\text{SO}_4)_3 \rightarrow \text{Ti}(\text{SO}_4)_2 + \text{FeSO}_4$
10. $\text{NaClO} + \text{H}_2\text{S} \rightarrow \text{NaCl} + \text{H}_2\text{SO}_4$
11. $\text{Fe}_3\text{O}_4 + \text{KMnO}_4 + \text{H}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + \text{K}_2\text{SO}_4 + \text{MnSO}_4 + \text{H}_2\text{O}$
12. $\text{NaCrO}_2 + \text{NaOH} + \text{H}_2\text{O}_2 \rightarrow \text{Na}_2\text{CrO}_4 + \text{H}_2\text{O}$
13. $\text{Ag} + \text{HNO}_3 \rightarrow \text{AgNO}_3 + \text{NO} + \text{H}_2\text{O}$
14. $\text{Ag}_2\text{S} + \text{HNO}_3 \rightarrow \text{AgNO}_3 + \text{NO} + \text{S} + \text{H}_2\text{O}$

(continued)



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David E. Newton



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