

QUESTION 2

(a) (i) The first ionization energy of an element is the amount of energy required to remove one electron from each atom in a mole of gaseous atoms, producing one mole of gaseous ions with a positive charge.

(ii) The electronic configuration of "Be" is $1s^2 2s^2$ and that of "B" is $1s^2 2s^2 2p^1$. All the subshells of Beryllium is filled and there is some special or extra energetic stability associated with that. This means that Beryllium is more stable than boron and hence has a higher ionization energy.

(b)

$$e \times F \equiv \frac{M}{It}$$

$$e \times 96500 \equiv \frac{48}{0.12 \times 500} \quad 0.015$$

Cross-multiplying

$$e = \frac{0.12 \times 500 \times 48}{96500 \times 0.015}$$

$$e = 1.99$$

$$e \approx 2$$

The charge is +2

(c) (i) Pure alumina in molten cryolite, Na_2AlF_6

- (ii) • Availability of raw materials
- Availability of electric power

(d) (i) Paramagnetism is a property exhibited by ferrous atoms or ions having one or more unpaired electrons which are attracted into a magnetic field. e.g. iron atom has four unpaired d $3d$ electrons.

(d) (ii) $2s^2$
I) $2s^2$

II) $2s^2$

III) $2s^2$

IV) $2s^2$

V) $2s^2$

VI) $2s^2$

VII) $2s^2$

VIII) $2s^2$

IX) $2s^2$

X) $2s^2$

XI) $2s^2$

XII) $2s^2$

XIII) $2s^2$

XIV) $2s^2$

XV) $2s^2$

XVI) $2s^2$

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JS Bruner



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 - The Rise of Digital Reading Answers Of Chemistry 2 Essay And 1 Obj
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers Of Chemistry 2 Essay And 1 Obj
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers Of Chemistry 2 Essay And 1 Obj
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers Of Chemistry 2 Essay And 1 Obj
 - Personalized Recommendations
 - Answers Of Chemistry 2 Essay And 1 Obj User Reviews and Ratings
 - Answers Of Chemistry 2 Essay And 1 Obj and Bestseller Lists
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 - Interactive Elements Answers Of Chemistry 2 Essay And 1 Obj
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 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers Of Chemistry 2 Essay And 1 Obj
9. Balancing eBooks and Physical Books Answers Of Chemistry 2 Essay And 1 Obj
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers Of Chemistry 2 Essay And 1 Obj
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answers Of Chemistry 2 Essay And 1 Obj
 - Setting Reading Goals Answers Of Chemistry 2 Essay And 1 Obj
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answers Of Chemistry 2 Essay And 1 Obj
 - Fact-Checking eBook Content of Answers Of Chemistry 2 Essay And 1 Obj
 - Distinguishing Credible Sources
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 - Utilizing eBooks for Skill Development
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