

## AP Chemistry Chapter 4 Practice Exam

D

1. The Claus reactions, shown below, are used to generate elemental sulfur from hydrogen sulfide.

How much sulfur (in grams) is produced from 21.5 grams of  $\text{O}_2$ ?

A) 97.0 g

B) 14.4 g

C) 21.5 g

D) 43.1 g

E) none of these

$$\frac{21.5 \text{ g O}_2}{32.00 \text{ g O}_2} \times \frac{1 \text{ mol O}_2}{1 \text{ mol O}_2} \times \frac{2 \text{ SO}_2}{3 \text{ O}_2} \times \frac{3 \text{ S}}{\text{SO}_2} \times \frac{32.06 \text{ g S}}{1 \text{ mol S}} = 43.1 \text{ g S}$$

B

2. Iron is produced from its ore by the reactions:

How many moles of  $\text{O}_2(g)$  are needed to produce 4.6 moles of  $\text{Fe}(s)$ ?A) 2.3 mole  $\text{O}_2$ B) 3.5 mole  $\text{O}_2$ C) 4.6 mole  $\text{O}_2$ D) 6.9 mole  $\text{O}_2$ E) 13.8 mole  $\text{O}_2$ 

$$\frac{4.6 \text{ mol Fe}(s)}{2 \text{ Fe}} \times \frac{3 \text{ CO}}{2 \text{ Fe}} \times \frac{1 \text{ O}_2}{2 \text{ CO}} = 3.45 \text{ mol O}_2$$

B

3. Suppose the reaction
- $\text{Ca}_3(\text{PO}_4)_2 + 3\text{H}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{H}_3\text{PO}_4$
- is carried out starting with 153 g of
- $\text{Ca}_3(\text{PO}_4)_2$

and 76.8 g of  $\text{H}_2\text{SO}_4$ . How much phosphoric acid will be produced?

A) 76.7 g

B) 31.1 g

C) 229.8 g

D) 115.1 g

E) 96.7 g

$$\frac{153 \text{ g Ca}_3(\text{PO}_4)_2}{310.18 \text{ g Ca}_3(\text{PO}_4)_2} \times \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} \times \frac{2 \text{ H}_3\text{PO}_4}{1 \text{ Ca}_3(\text{PO}_4)_2} = 0.97 \text{ mol H}_3\text{PO}_4$$

$$\frac{76.8 \text{ g H}_2\text{SO}_4}{98.08 \text{ g H}_2\text{SO}_4} \times \frac{1 \text{ mol H}_2\text{SO}_4}{1 \text{ mol H}_2\text{SO}_4} \times \frac{2 \text{ H}_3\text{PO}_4}{3 \text{ H}_2\text{SO}_4} = 0.51 \text{ mol H}_3\text{PO}_4$$

B

- 4.
- $\text{SO}_2$
- reacts with
- $\text{H}_2\text{S}$
- as follows:

When 7.50 g of  $\text{H}_2\text{S}$  reacts with 12.75 g of  $\text{SO}_2$ , which statement applies?

A) 6.38 g of sulfur are formed.

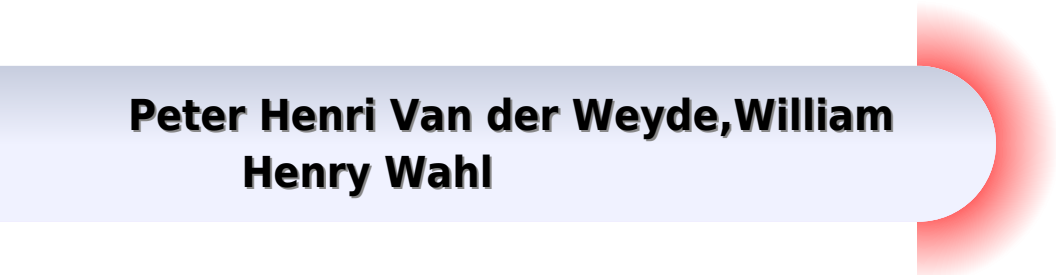
B) 10.6 g of sulfur are formed.C) 0.0216 moles of  $\text{H}_2\text{S}$  remain.D) 1.13 g of  $\text{H}_2\text{S}$  remain.E)  $\text{SO}_2$  is the limiting reagent.

$$\frac{7.50 \text{ g H}_2\text{S}}{34.03 \text{ g H}_2\text{S}} \times \frac{1 \text{ mol H}_2\text{S}}{1 \text{ mol H}_2\text{S}} \times \frac{3 \text{ S}}{2 \text{ H}_2\text{S}} \times \frac{32.06 \text{ g S}}{1 \text{ mol S}} = 10.89 \text{ g S}$$

$$\frac{12.75 \text{ g SO}_2}{64.06 \text{ g SO}_2} \times \frac{1 \text{ mol SO}_2}{1 \text{ mol SO}_2} \times \frac{3 \text{ S}}{1 \text{ SO}_2} \times \frac{32.06 \text{ g S}}{1 \text{ mol S}} = 19.13 \text{ g S}$$

# Answers For Ap Chemistry Test Open Ended

**Peter Henri Van der Weyde, William  
Henry Wahl**



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*Learning and Understanding* National Research Council, Division of Behavioral and Social Sciences and Education, Center for Education, Committee on Programs for Advanced Study of Mathematics and Science in American High Schools, 2002-08-06 This book takes a fresh look at programs for advanced studies for high school students in the United States with a particular focus on the Advanced Placement and the International Baccalaureate programs and asks how advanced studies can be significantly improved in general It also examines two of the core issues surrounding these programs they can have a profound impact on other components of the education system and participation in the programs has become key to admission at selective institutions of higher education By looking at what could enhance the quality of high school advanced study programs as well as what precedes and comes after these programs this report provides teachers parents curriculum developers administrators college science and mathematics faculty and the educational research community with a detailed assessment that can be used to guide change within advanced study programs Children's Reading Comprehension and Assessment Scott G. Paris, Steven A. Stahl, 2005-03-23 The volume identifies critical issues and cutting edge research in assessing children's reading comprehension Problems and Problem Solving in Chemistry Education Georgios Tsaparlis, 2021-05-17 Problem solving is central to the teaching and learning of chemistry at secondary tertiary and post tertiary levels of education opening to students and professional chemists alike a whole new world for analysing data looking for patterns and making deductions As an important higher order thinking skill problem solving also constitutes a major research field in science education Relevant education research is an ongoing process with recent developments occurring not only in the area of quantitative computational problems but also in qualitative problem solving The following situations are considered some general others with a focus on specific areas of chemistry quantitative problems qualitative reasoning metacognition and resource activation deconstructing the problem solving process an overview of the working memory hypothesis reasoning with the electron pushing formalism scaffolding organic synthesis skills spectroscopy for structural characterization in organic chemistry enzyme kinetics problem solving in the academic chemistry laboratory chemistry problem solving in context team based active learning technology for molecular representations IR spectra simulation and computational quantum chemistry tools The book concludes with methodological and epistemological issues in problem solving research and other perspectives in problem solving in chemistry With a foreword by George Bodner **Barron's AP Statistics** Martin Sternstein, 1998 This all new prep manual reflects the growing trend in statistics courses taken by high school students preparing for college Recommended as a supplement to classroom textbooks or as a self teaching preparation this guide offers 545 practice questions and solutions including five full length practice exams Topics reviewed include graphical displays distributions bivariate data and more **Chemistry** Steven S. Zumdahl, 1993 *Journal of Research and Development in Education*, 1986 Applied Science & Technology Index, 2000 **The Education Index**

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