

Gravimetric Analysis of Chloride in Solution

Experiment 10
Lab 10
Date: _____

Introduction:

This experiment involves the gravimetric analysis of a solution containing chloride ions. The principle of gravimetric analysis is based on the fact that the mass of a substance is proportional to the number of moles of the substance. In this experiment, the chloride ions in the solution are precipitated as silver chloride (AgCl) by adding a known volume of a silver nitrate solution. The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution. The experiment is performed in a fume hood to avoid exposure to silver nitrate, which is a strong oxidizing agent and can cause skin irritation. The experiment is also performed in a dry environment to avoid the formation of silver chloride hydrate, which can lead to inaccurate results.

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Procedure:

The procedure for this experiment is as follows: 1. Weigh a known mass of the sample (e.g., 0.1000 g) and transfer it to a beaker. 2. Add a known volume of water (e.g., 100 mL) to the beaker. 3. Add a known volume of silver nitrate solution (e.g., 10 mL) to the beaker. 4. Stir the mixture for 10 minutes. 5. Filter the mixture through a filter paper. 6. Wash the filter paper with distilled water. 7. Dry the filter paper in a desiccator. 8. Weigh the dried filter paper. 9. Calculate the mass of chloride ions in the sample.

Data Analysis:

	Sample 1	Sample 2
Mass of sample	0.1000 g	0.1000 g
Mass of precipitate	0.1500 g	0.1500 g
Mass of chloride ion	0.0500 g	0.0500 g

The mass of chloride ion in the sample is calculated as follows: $m_{\text{Cl}^-} = m_{\text{AgCl}} \times \frac{m_{\text{Cl}^-}}{m_{\text{AgCl}}}$

For Sample 1: $m_{\text{Cl}^-} = 0.1500 \text{ g} \times \frac{35.45}{143.32} = 0.0375 \text{ g}$

For Sample 2: $m_{\text{Cl}^-} = 0.1500 \text{ g} \times \frac{35.45}{143.32} = 0.0375 \text{ g}$

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