

CHM 111 Analysis of Aluminum-Zinc Alloy using Ideal Gas Law

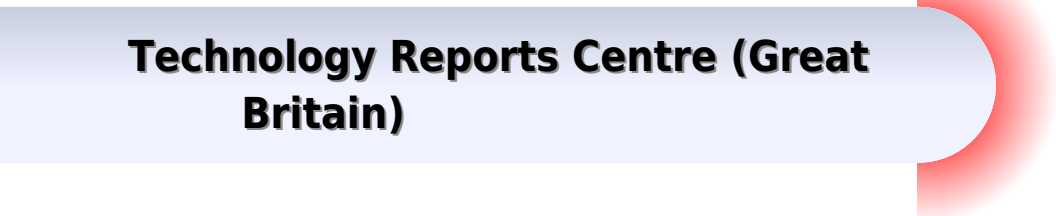
Report Sheet

Data Sheet:

	Trial 1	Trial 2
Mass of gelatin capsule	<u>1.7382</u> g	<u> </u> g
Mass of alloy sample plus capsule	<u>1.9852</u> g	<u> </u> g
Mass of empty beaker	<u>104.3024</u> g	<u> </u> g
Mass of beaker plus displaced water	<u>346.225</u> g	<u> </u> g
Barometric pressure	<u> </u> mm Hg	<u> </u> mm Hg
Temperature	<u>24</u> °C	<u> </u> °C
Calculations Mass of alloy sample	<u>0.247</u> g	<u> </u> g
Mass of displaced water	<u>241.9226</u> g	<u> </u> g
Volume of displaced water	<u> </u> mL	<u> </u> mL
Volume of H ₂ , V	<u> </u> L	<u> </u> L
Temperature of H ₂ , T	<u>297.15</u> K	<u> </u> K
Vapor pressure of water at T	<u>22.4</u> mm Hg	<u> </u> mm Hg
Pressure of dry H ₂ ,	<u>755.7</u> mm Hg; <u>0.9943</u> atm	<u> </u> mm Hg; <u> </u> atm
Moles H ₂ from sample,	<u> </u> moles	<u> </u> moles
Moles H ₂ per gram of sample,	<u> </u> moles/g	<u> </u> moles/g %
Al (read from graph)	<u> </u> %	<u> </u> %
Unknown no. <u>3</u>		

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